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# Water-Soluble Anthraquinone Photocatalysts Enable Methanol-Driven Enzymatic Halogenation and Hydroxylation Reactions

Bo Yuan, Durga Mahor, Qiang Fei, Ron Wever, Miguel Alcalde, Wuyuan Zhang,\* and Frank Hollmann\*

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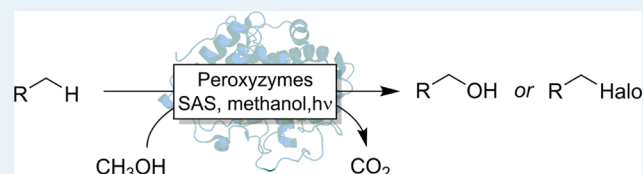
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**ABSTRACT:** Peroxozymes simply use  $\text{H}_2\text{O}_2$  as a cosubstrate to oxidize a broad range of inert C–H bonds. The lability of many peroxozymes against  $\text{H}_2\text{O}_2$  can be addressed by a controlled supply of  $\text{H}_2\text{O}_2$ , ideally in situ. Here, we report a simple, robust, and water-soluble anthraquinone sulfonate (SAS) as a promising organophotocatalyst to drive both haloperoxidase-catalyzed halogenation and peroxxygenase-catalyzed oxyfunctionalization reactions. Simple alcohols, methanol in particular, can be used both as a cosolvent and an electron donor for  $\text{H}_2\text{O}_2$  generation. Very promising turnover numbers for the biocatalysts of up to 318 000 have been achieved.

**KEYWORDS:** photobiocatalysis, peroxozymes, halogenation, hydroxylation,  $\text{H}_2\text{O}_2$  generation



## INTRODUCTION

Peroxozymes (i.e., oxidative enzymes utilizing  $\text{H}_2\text{O}_2$  as a stoichiometric oxidant) are on the rise. As biocatalysts, they entail the benefits associated with enzymes such as operating under mild (i.e., energy-saving) reaction conditions and often exhibiting superb regio-, chemo-, and enantio-selectivity. Compared to other oxidizing enzymes such as P450 monooxygenases, they excel by their simplicity in terms of oxygen activation.<sup>1</sup> Instead of relying on complex enzyme cascades for the activation of molecular oxygen, peroxozymes use already activated oxygen in the form of  $\text{H}_2\text{O}_2$ .<sup>2</sup> Nevertheless, in most cases, simple bulk addition of all  $\text{H}_2\text{O}_2$  is not advisable because of undesired side reactions and/or limited stability of biocatalysts against  $\text{H}_2\text{O}_2$ . Slow, continuous, or portion-wise supply of  $\text{H}_2\text{O}_2$  is one alternative to maintain overall low concentrations of  $\text{H}_2\text{O}_2$  but leads to significant dilution effects. Also, so-called “hot spots” of locally high  $\text{H}_2\text{O}_2$  concentrations at the entry point of the stock solution into the reaction medium eventually inactivate the biocatalysts. Finally, especially in the case of hypohalite-forming enzymes (vide infra), excess  $\text{H}_2\text{O}_2$  is wasted in a futile disproportionation reaction yielding  $\text{H}_2\text{O}$  and  $^1\text{O}_2$ .<sup>3,4</sup> The latter may also be expected to lead to undesired side reactions. Therefore, in situ generation of  $\text{H}_2\text{O}_2$  from  $\text{O}_2$  has emerged as a viable alternative.<sup>2</sup> A broad range of enzymatic, electrochemical, and chemical in situ  $\text{H}_2\text{O}_2$  generation systems have been proposed in the past few years. Photochemical approaches for the in situ reduction of  $\text{O}_2$  to  $\text{H}_2\text{O}_2$  are attractive because they also enable using sacrificial electron donors such as EDTA<sup>5,6</sup> or simple alcohols.<sup>7</sup> Enzymatic  $\text{H}_2\text{O}_2$  generation methods are more restricted in terms of sacrificial electron donors that can be used.<sup>8–10</sup> Furthermore, combining photocatalysis with biocatalysis offers new possibilities for organic synthesis ranging from new regeneration approaches for

cofactor-dependent enzymes,<sup>11–25</sup> photoenzymatic cascades,<sup>26</sup> and “new to nature” reactions catalyzed by photoexcited enzymes.<sup>27–32</sup> Yet, the combination of photocatalysis with biocatalysis is not always unproblematic due to issues of photobleaching<sup>33</sup> and formation of reactive oxygen species.<sup>34–36</sup>

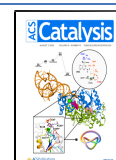
Therefore, we became interested in the photocatalyst sodium anthraquinone sulfonate (SAS) for the oxidation of small sacrificial electron donors such as methanol<sup>37,38</sup> and reductive activation of ambient  $\text{O}_2$  to  $\text{H}_2\text{O}_2$ . Compared to established heterogeneous photocatalysts,<sup>7,39</sup> SAS is homogeneously dissolved in the reaction medium, which may alleviate the sluggish reaction kinetics by eliminating diffusion limitations of the heterogeneous photocatalysts. Furthermore, the mechanism of anthraquinone-mediated oxygen activation<sup>40</sup> does not involve long-lived radical species, which in previous studies have been observed to impair the stability of the biocatalysts used. SAS has also been studied extensively as a photocatalyst for aerobic oxidation of a broad range of different alcohols using molecular oxygen as a stoichiometric electron acceptor (yielding  $\text{H}_2\text{O}_2$  as byproduct).<sup>37,41–43</sup>

We, therefore, decided to evaluate SAS-catalyzed oxidation of simple alcohols to promote peroxozyme-catalyzed oxidation reactions. As the first model enzyme, we used vanadium-dependent chloroperoxidase from *Curvularia inaequalis* (CiV-CPO).<sup>4,34,44–53</sup> CiV-CPO oxidizes  $\text{Cl}^-$ ,  $\text{Br}^-$ , and  $\text{I}^-$  to the corresponding hypohalites, which then undergo spontaneous

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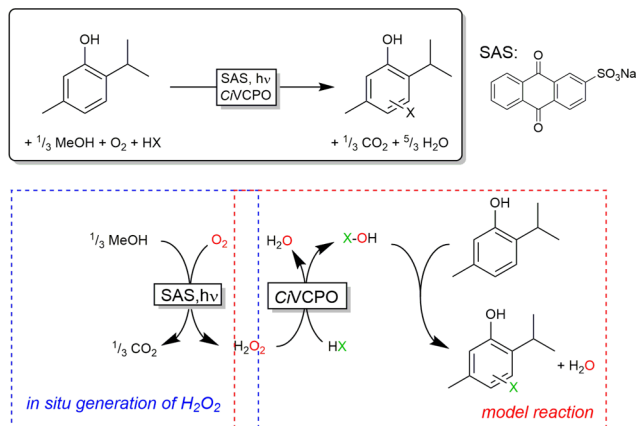
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chemical oxidation and halogenation reactions. We envisioned a productive coupling of SAS-catalyzed  $\text{H}_2\text{O}_2$  generation with  $\text{CiVCPO}$ -induced halogenation of thymol (Scheme 1).

**Scheme 1. Photoenzymatic Halogenation Combining Photocatalytic In Situ Generation of  $\text{H}_2\text{O}_2$  Via  $\text{O}_2$  Reduction in the Presence of Methanol to Drive  $\text{CiVCPO}$ -Initiated Halogenation of Thymol<sup>a</sup>**



<sup>a</sup>Upper: overall reaction, lower: dissection into photochemical  $\text{H}_2\text{O}_2$  generation (blue) and chemoenzymatic halogenation of thymol (red).

## RESULTS AND DISCUSSION

For the first set of experiments, we chose methanol as a sacrificial electron donor. Due to the excellent solvent stability of the biocatalyst, we used 40% (v/v) of methanol in phosphate buffer (pH 6). Hence, methanol served two purposes, as a sacrificial electron donor and as a cosolvent for the hydrophobic thymol starting material. Illuminating SAS alone with the visible light in the presence of methanol and ambient air led to an accumulation of  $\text{H}_2\text{O}_2$  in the reaction mixture (Figure S1). After approximately 5 h (ca. 2 mM  $\text{H}_2\text{O}_2$ ), the  $\text{H}_2\text{O}_2$  accumulation ceased, most likely reaching a steady state between SAS-mediated  $\text{H}_2\text{O}_2$  formation and decomposition (vide infra).

Performing the same experiment albeit in the presence of  $\text{CiVCPO}$  and thymol, a linear product accumulated over more than 30 h, reaching more than 95% conversion of the starting material. Overall, 9.14 mM 4-bromothymol (**1b**) was produced with traces of the 2-bromo isomer (**1a**) and the dibromination product (2,4-dibromo thymol, **1c** stemming from the sequential halogenation of the primary product **1a**)<sup>49,54</sup> (Figure 1a). On performing the experiment either in the darkness or in the absence of SAS (under otherwise identical conditions), no conversion of the starting material was observed. In the absence of  $\text{CiVCPO}$ , traces of 4-bromothymol were observed upon prolonged reaction times. We, hence, conclude that the reaction indeed proceeds via the sequence outlined in Scheme 1. Substitution of the phosphate buffer with Tris-HCl gave almost identical results (Table S1). Interestingly, under otherwise identical conditions, performing the experiment in sodium citrate buffer gave only poor results. Currently, we lack a plausible explanation for this observation.

Increasing the photocatalyst concentration from 0.25 to 2 mM increased the product (**1b**) formation rate from 0.48 to 0.66 mM  $\text{h}^{-1}$  (Figure 1b). The nonproportional increase of **1b** may indicate a decreasing optical transparency of the reaction mixture and a resulting suboptimal illumination of the entire

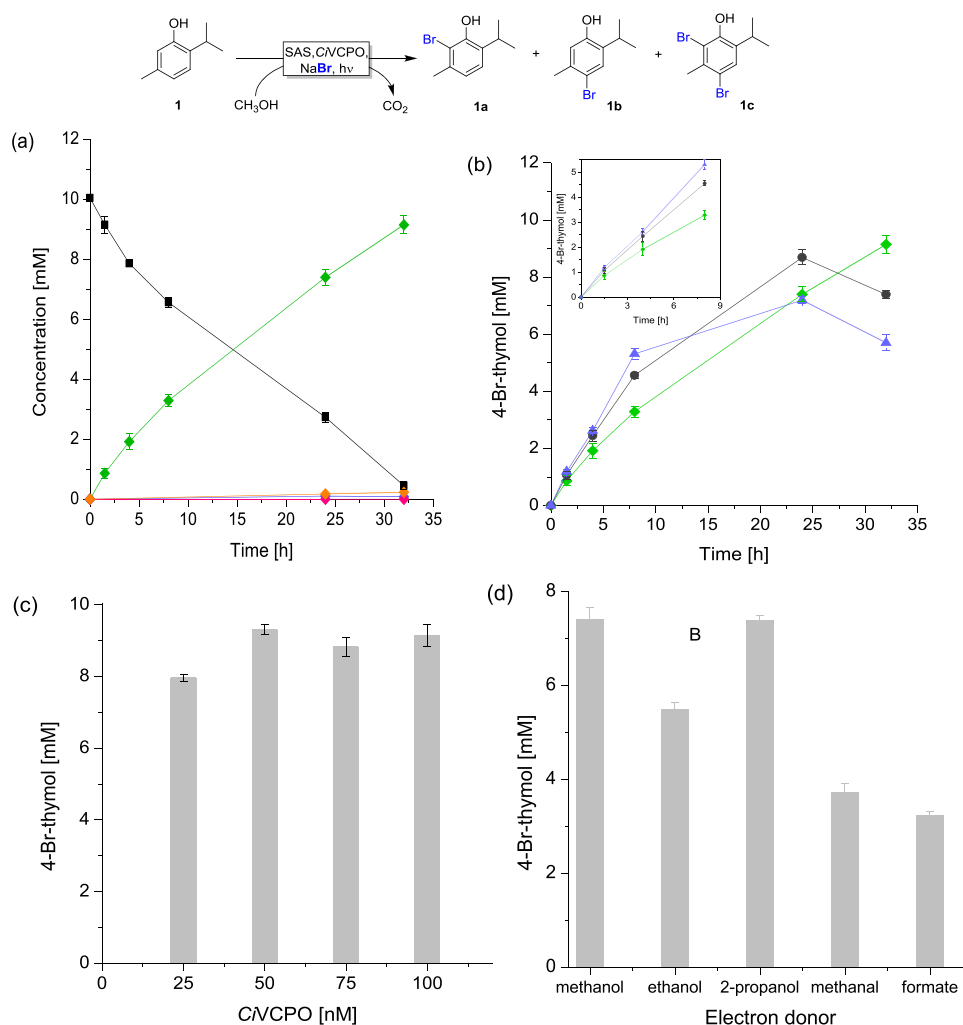
reaction mixture, and hence a suboptimal utilization of the photocatalysts. Similar effects have been observed in other photocatalytic systems such as algal fermentation. Increasing the photocatalyst concentration, however, also negatively influenced the overall selectivity of the reaction and significant amounts of the dibromination product (**1c**) accumulated in the reaction mixture. In general, upon prolonged reaction times (i.e., upon near-full conversion of the thymol starting material), **1b** was further brominated to **1c** (Figure S2). Similarly, using **1b** as a starting material also led to the formation of **1c**. The biocatalyst concentration had no significant influence on the overall product formation (Figure 1c). Together with the previous observation, this may indicate that  $\text{H}_2\text{O}_2$  generation reaction was overall rate limiting. This assumption is supported by the almost linear correlation observed between the product formed and the intensity of the light source (Figure S3). Also, the wavelength had a significant influence on the overall productivity (Figure S4). Applying blue light gave higher product concentrations as compared to bathochrome green or red light, which we attribute to the different overlaps with the SAS absorption spectrum (Figure S5).

Next, we turned our attention to the concentration and nature of the sacrificial electron donor. Increasing the methanol concentration from 10% (v/v) to 30% (v/v) also resulted in an increased product accumulation (Figure S6). Further increase of the methanol concentration up to 70% (v/v) did not significantly influence the product formation (rate).

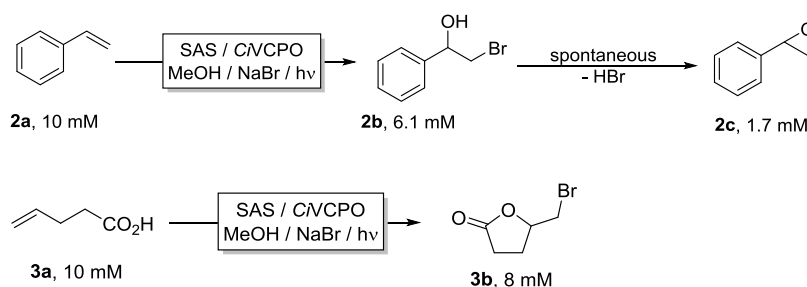
The system does not necessarily rely on methanol as also other alcohols such as ethanol and 2-propanol yielded promising results (Figure 1d). Noteworthy, the methanol oxidation products formaldehyde and formate were also oxidized at significant rates yielding  $\text{H}_2\text{O}_2$ , pointing toward triple oxidation of methanol to  $\text{CO}_2$ . It is also possible to scale up the halogenation reaction to 100 mL, in which 967 mg (45% isolated yield) of halogenated thymol was obtained (for details, see the scaling-up experiments in the Supporting Information).

The synthetic scope of the new photoenzymatic reaction system was essentially identical to the traditional application of  $\text{CiVCPO}$  using stoichiometric  $\text{H}_2\text{O}_2$  (Figures 2 and S7). Similar to previous reports,<sup>47</sup> the hydroxybromination of styrene proceeded smoothly to the desired hydroxybromide (**2b**, 61%) and epoxide (**2c**, 17% originating from spontaneous oxirane formation). Also, the recently reported bromolactonization of 4-pentenoic acid gave the desired bromolactone (**3b**) in 80% yield.<sup>44,46</sup>

Overall, these experiments have demonstrated that  $\text{CiVCPO}$ -initiated bromofunctionalization reactions can be driven by photocatalytic  $\text{H}_2\text{O}_2$  generation using SAS as a photocatalyst. In these experiments,  $\text{CiVCPO}$  performed up to 318 000 catalytic turnovers (on average  $2.7 \text{ s}^{-1}$  over 32 h). Compared to previous systems coupling  $\text{CiVCPO}$  to other heterogeneous  $\text{H}_2\text{O}_2$ -generating photocatalysts,<sup>34</sup> the present system enabled significantly more robust production schemes (i.e., continuous product formation over at least 30 h). As  $\text{CiVCPO}$  itself is known to be highly robust against  $\text{H}_2\text{O}_2$ , we believe that this observation is in line with the hypothesized absence of reactive oxygen species such as hydroxyl radicals or superoxide.<sup>34</sup> Interestingly enough, under the current conditions, bulk addition of  $\text{H}_2\text{O}_2$  gave higher catalyst turnover numbers (Table S2). We attribute this to the insufficient light penetration into the reaction mixture and therefore suboptimal utilization of the photocatalyst. Future optimization will focus on alternative



**Figure 1.** Halogenation of thymol by combining CiVCPO and visible light-driven in situ generation of  $\text{H}_2\text{O}_2$  using SAS. (a) Conversion of thymol (1, ■) into 4-brominated thymol (1b, ♦) in the presence of CiVCPO (100 nM) and SAS (0.5 mM), and control reactions in the dark (♦) in the absence of CiVCPO (♦) or SAS (♦). (b) Influence of varied concentrations of SAS (♦ = 0.25 mM, ● = 1 mM, and ▲ = 2 mM) and (c) CiVCPO on the reaction course. (d) Other cosolvent (as well as electron donor) investigated. Reaction conditions were as follows: [substrate] = 10 mM, [CiVCPO] = 25–100 nM, [SAS] = 0.5–2 mM, [NaBr] = 25 mM, pH 6.0 (NaPi buffer, 60 mM), 40% of cosolvent, and visible light illumination ( $\lambda > 400$  nm). The concentration of methanol and formate was 100 mM. The yielded products were quantified by gas chromatography. Error bars represent the standard deviation of duplicate experiments.



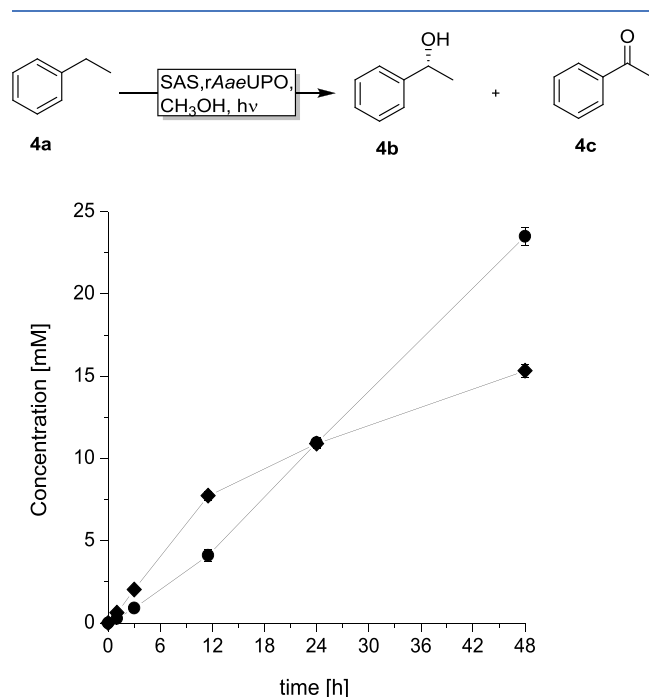
**Figure 2.** Hydroxybromination and bomocyclization reactions by combining CiVCPO and visible light-driven in situ generation of  $\text{H}_2\text{O}_2$  using SAS. Reaction conditions were as follows: [substrate] = 10 mM, [CiVCPO] = 50 nM, [SAS] = 0.5 mM, pH 6.0 (NaPi buffer, 60 mM), 40% of methanol, 32 h, and visible light illumination ( $\lambda > 400$  nm). Experiments were performed in independent duplicates.

reactor concepts such as flow chemistry setups or internal illumination to alleviate this limitation.

To investigate whether somewhat less robust peroxozymes can profit from the seemingly milder in situ  $\text{H}_2\text{O}_2$  generation system based on SAS, we turned our attention to the recombinant, evolved peroxxygenase from *Agrocye aegeri-*

*ta* (rAaeUPO).<sup>55–58</sup> From a synthetic point of view, rAaeUPO is very interesting as it catalyzes the specific oxyfunctionalization of a broad range of compounds.<sup>7,10,55–57,59–69</sup> As a heme-dependent enzyme, it however, is also prone to irreversible oxidative inactivation by  $\text{H}_2\text{O}_2$  or other reactive oxygen species. Therefore, we next performed an experiment on the hydrox-

ylation of ethylbenzene using *rAaeUPO* and SAS-mediated  $\text{H}_2\text{O}_2$  generation (Figure 3).



**Figure 3.** Hydroxylation of ethylbenzene into (*R*)-1-phenyl ethanol (◆) and acetophenone (●) by combining *rAaeUPO* and visible light-driven in situ generation of  $\text{H}_2\text{O}_2$  using SAS. Reaction conditions were as follows: [substrate] = 50 mM, [*rAaeUPO*] = 100 nM, [SAS] = 0.5 mM, 40% of methanol, and visible light illumination ( $\lambda > 400$  nm). Experiments were performed as independent duplicates.


Pleasingly, the steady accumulation of the ethylbenzene oxyfunctionalization products ((*R*)-1-phenyl ethanol, **4b** and acetophenone, **4c**) was observed for at least 48 h, indicating a very high robustness of *rAaeUPO* under the reaction conditions chosen. In contrast to previous experiments on the *rAaeUPO*-catalyzed hydroxylation of ethylbenzene, we observed a low enantio-selectivity (34% ee) and a high degree of overoxidation, which can be attributed to SAS-catalyzed oxidation of the substrate **4a** into **4b**, and the enzymatic product **4b** into **4c**. Further optimization of the ratio of SAS to *rAaeUPO* alleviates this issue.


Again, we compared the performance of the biocatalyst using the proposed photochemical  $\text{H}_2\text{O}_2$  generation system with bulk addition of  $\text{H}_2\text{O}_2$  (Figure S8 and Table S2). In contrast to  $\text{H}_2\text{O}_2$ -tolerant CiVCPO, *rAaeUPO* performed significantly better under in situ  $\text{H}_2\text{O}_2$  generation.


While overoxidation is undesired in the case of the stereoselective hydroxylation of ethylbenzene, it is desirable in the case of the oxidation of cyclohexane to cyclohexanone as the precursor for  $\epsilon$ -caprolactam.<sup>70–75</sup> As *rAaeUPO* has previously been reported to hydroxylate (cyclo)alkanes, we envisioned a photoenzymatic cascade to transform cyclohexane into the corresponding alcohol or ketone.<sup>76</sup> In a first experiment, we used cyclohexane solubilized in an aqueous reaction buffer and the cosubstrate methanol in the presence of *rAaeUPO* (Table 1). Cyclohexane was converted at reasonable conversion (71%), yielding an approximate 6:1 ratio of cyclohexanone to cyclohexanol, which we attribute to the photocatalytic, SAS-mediated oxidation of cyclohexanol. This ratio, however, can be

**Table 1.** Photoenzymatic Hydroxylation of Cyclohexane in a Two-Phase Reaction System<sup>a</sup>

organic phase (substrate reservoir & product sink)


  
**5a**

  
**5b**


  
**5c**

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aqueous reactive phase

  
 $\text{CH}_3\text{OH}$

**rAaeUPO,  
SAS,  $h\nu$**

  
 $\text{CO}_2$

	5b [mM]	5c [mM]	TON (SAS)	TON (rAaeUPO)
monophase	2.6	15.1	35	177 000 (71% conv.)
two-phase	13.3	3.5	33	168 000

<sup>a</sup>Reaction conditions: substrate/NaPi buffer (pH 6.0, 60 mM) phase ratio = 1:1 (v/v). In the aqueous phase: [*rAaeUPO*] = 100 nM, [SAS] = 0.5 mM, 40% of methanol, 32 h, visible light illumination ( $\lambda > 400$  nm). For the monophase reaction, 25 mM of substrate was added.

inverted using cyclohexane as a second phase under otherwise identical conditions (Table 1). Here, a significant proportion of the primarily formed cyclohexanol partitioned into the (catalyst-free) organic phase and thereby was not available for further SAS- or *rAaeUPO*-catalyzed further oxidation.

In both cases, very good catalytic turnover of the biocatalysts above 150 000 and reasonable turnover numbers in the range of 35 for the photocatalyst were calculated.

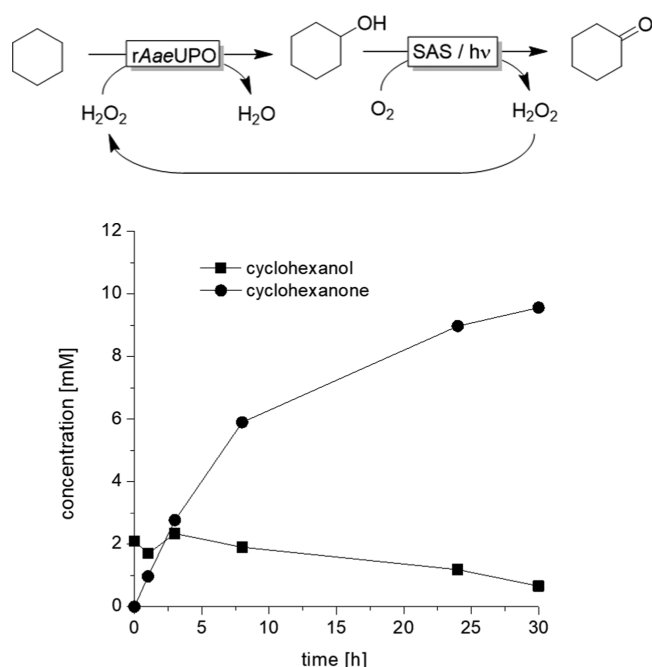
The SAS-catalyzed aerobic “overoxidation” of cyclohexanol to cyclohexanone yields  $\text{H}_2\text{O}_2$  as a byproduct, which itself can be used to promote the *rAaeUPO*-catalyzed hydroxylation of cyclohexane. This way, an overall aerobic oxidation of cyclohexane to cyclohexanone can be imagined (Figure 4). In a first experiment, we used 2 mM cyclohexanol to “kick-start” the photocatalytic generation of  $\text{H}_2\text{O}_2$  and therewith the entire cascade.

## CONCLUSIONS

In this work, we have expanded the scope of photogeneration of  $\text{H}_2\text{O}_2$  to promote peroxzyme-catalyzed halogenation and hydroxylation reactions. Water-soluble sodium anthraquinone sulfonate is a promising alternative to established heterogeneous photocatalysts. In contrast to the latter, SAS enables highly robust peroxzyme reactions, and the turnover numbers for the two peroxzymes used here (CiVCPO and *rAaeUPO*) reached 318 000 and 177 000, respectively. Most likely, this is due to the avoidance of oxygen radical species in the case of SAS-catalyzed  $\text{H}_2\text{O}_2$  generation.

Furthermore, SAS also functions as an oxidation catalyst for peroxxygenase-derived conversion of alcohol into ketones in a simple and self-sufficient manner. The catalytic activity of SAS itself is yet not optimal. Especially, the optical intransparency of the reaction mixtures suggests that light does not deeply penetrate the reaction mixtures. As a consequence, the majority of SAS present in the reaction mixtures is not illuminated and therefore remains “idle.” Further optimization will therefore





**Figure 4.** Self-sufficient, aerobic oxidation of cyclohexane to cyclohexanone. Reaction conditions were as follows: [cyclohexanol] = 2 mM, [cyclohexane] = 25 mM, [rAaeUPO] = 100 nM, [SAS] = 0.5 mM in NaPi buffer (pH 6.0, 60 mM). The values shown stem from one experiment (no duplicates).

focus on alternative reactor concepts with optimized surface to bulk ratios enabling more efficient utilization of the photocatalyst.

## EXPERIMENT SECTION

**Catalyst Preparation.** The heterologous expression and purification of vanadium chloroperoxidase from *C. inaequalis* (CiVCPO) were performed according to reported procedures.<sup>77</sup> The recombinant unspecific peroxygenase from *A. aegerita* (rAaeUPO) was produced and purified by following previous methods.<sup>57</sup> The photocatalyst SAS was purchased from Sigma-Aldrich and used without extra treatment.

**Photoenzymatic Halogenation Reactions.** The photochemical enzymatic halogenation reactions using CiVCPO were performed at 30 °C in 1.0 mL of sodium phosphate buffer (NaPi, pH 6.0, 60 mM). Specifically, a stock solution (5 mM) of sodium anthraquinone sulfonate (SAS) in above NaPi buffer and thymol (100 mM) in methanol were first prepared. Hundred microliters of each stock solution was added to 795  $\mu\text{L}$  of a premixed solution (300  $\mu\text{L}$  methanol and 495  $\mu\text{L}$  NaPi buffer) in a 4 mL glass vial. Afterward, CiVCPO was added (5  $\mu\text{L}$ ). In the final solution, the reaction conditions were as follows: [substrate] = 10 mM, [CiVCPO] = 50 nM, [NaBr] = 25 mM, [SAS] = 0.5 mM, pH 6.0 (NaPi buffer, 60 mM), and 40% of methanol in 1.0 mL. As the final step, the reaction vial was closed and exposed to visible light (Philips 7748XHP 150 W, white light bulb) under gentle magnetic stirring (200 rpm). At intervals, aliquots were withdrawn, extracted with ethyl acetate (containing 5 mM dodecane as an internal reference, extraction ratio 1:2), and analyzed by gas chromatography (SHIMADZU). All the above reactions were performed in independent duplicates.

**Photoenzymatic Hydroxylation Reactions.** The photochemical enzymatic hydroxylation reactions using rAaeUPO were performed using a very similar approach, as described for

halogenation reactions. In a typical monophase reaction, a stock solution (5 mM) of sodium anthraquinone sulfonate (SAS) in NaPi buffer (pH 6.0, 60 mM) was first prepared. Hundred microliters of the stock solution was added to 895  $\mu\text{L}$  of a premixed solution (400  $\mu\text{L}$  methanol and 495  $\mu\text{L}$  NaPi buffer) in a 4 mL glass vial. Afterward, the substrate (ethylbenzene or cyclohexane) rAaeUPO was added. The final reaction conditions were as follows: [substrate] = 50 mM, [rAaeUPO] = 100 nM, [SAS] = 0.5 mM, pH 6.0 (NaPi buffer, 60 mM), and 40% of methanol in 1.0 mL at 30 °C. The reaction vial was closed and exposed to visible light under gentle magnetic stirring (200 rpm).

In a typical two-phase reaction, approximately 500  $\mu\text{L}$  of a premixed solution (200  $\mu\text{L}$  methanol and 300  $\mu\text{L}$  NaPi buffer) was added in a 4 mL glass vial. rAaeUPO was first added to this mixture, followed by addition of 500  $\mu\text{L}$  of substrate (ethylbenzene or cyclohexane) as the organic phase. In the final solution, the reaction conditions were as follows: substrate/NaPi buffer (pH 6.0, 60 mM) phase ratio = 1:1 (v/v). In the aqueous phase: [rAaeUPO] = 100 nM, [SAS] = 0.5 mM, 40% of methanol, 30 °C. All the above photoenzymatic hydroxylation reactions were performed in independent duplicates.

**Scaling Up the Synthesis of 4-Br-Thymol.** The scale-up reactions were performed in 100 mL for each batch (six batches in total). The reaction conditions were as follows: [substrate] = 15 mM, [CiVCPO] = 75 nM, [NaBr] = 25 mM, [SAS] = 0.5 mM, pH 6.0 (NaPi buffer, 60 mM), 40% methanol, and 30 °C. The mixture in a transparent DURAN bottle was placed (six batches at the same time) and stirred gently. The mixture was irradiated for 36 h at 30 °C under visible light. At the end of the reaction, the organic compounds were extracted using ethyl acetate (3 $\times$ ). The organic phase was combined and dried over anhydrous  $\text{Na}_2\text{SO}_4$ . After evaporating the ethyl acetate under reduced pressure, a yellowish oil was obtained. The crude product was purified by silica column using petroleum ether/ethyl acetate (40:2, v/v) as an eluent.

## ASSOCIATED CONTENT

### Supporting Information

The Supporting Information is available free of charge at <https://pubs.acs.org/doi/10.1021/acscatal.0c01958>.

Preparation of the enzymes, the photoenzymatic reaction setup, analytical data, and additional results (PDF)

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## Notes

The authors declare no competing financial interest.

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