

The Bi³⁺ 6s and 6p electron binding energies in relation to the chemical environment of inorganic compounds

Awater, Roy H.P.; Dorenbos, Pieter

DOI

[10.1016/j.jlumin.2016.12.021](https://doi.org/10.1016/j.jlumin.2016.12.021)

Publication date

2017

Document Version

Accepted author manuscript

Published in

Journal of Luminescence

Citation (APA)

Awater, R. H. P., & Dorenbos, P. (2017). The Bi³⁺ 6s and 6p electron binding energies in relation to the chemical environment of inorganic compounds. *Journal of Luminescence*, 184, 221-231.
<https://doi.org/10.1016/j.jlumin.2016.12.021>

Important note

To cite this publication, please use the final published version (if applicable).
Please check the document version above.

Copyright

Other than for strictly personal use, it is not permitted to download, forward or distribute the text or part of it, without the consent of the author(s) and/or copyright holder(s), unless the work is under an open content license such as Creative Commons.

Takedown policy

Please contact us and provide details if you believe this document breaches copyrights.
We will remove access to the work immediately and investigate your claim.

Vacuum Referred Binding Energies of Bi^{3+} in Insulators Based on the Metal-to-Metal Charge Transfer Energy

Roy H. P. Awater* and Pieter Dorenbos

*Luminescence Materials Research Group (FAME-LMR), Department of Radiation Science
and Technology, Faculty of Applied Sciences, Delft University of Technology, 2629 JB
Delft, The Netherlands*

E-mail: R.H.P.Awater@tudelft.nl

Abstract

Introduction

The luminescence of the Bi^{3+} activator ion in a variety of host compounds has been extensively studied over the last 50 years.¹ The Bi^{3+} ion has a $6s^2$ outer electron configuration with the $^1\text{S}_0$ ground state. Optical transitions to the $6s^16p^1$ configuration result in the $^3\text{P}_{0,1,2}$ triplet and $^1\text{P}_1$ singlet excited states (in order of increasing energy). The optical transitions from the $^1\text{S}_0$ ground state to the $^3\text{P}_1$, $^3\text{P}_2$ and $^1\text{P}_1$ excited states are labeled A, B and C, respectively (see Fig. 1). The $^1\text{S}_0 \rightarrow ^3\text{P}_0$ and $^1\text{S}_0 \rightarrow ^3\text{P}_2$ are spin-forbidden, although the transition to the $^3\text{P}_2$ can be induced by coupling with unsymmetrical lattice vibrational modes.² As a result of spin-orbit coupling and mixing with the $^1\text{P}_1$ state, the $^1\text{S}_0 \rightarrow ^3\text{P}_1$ transition becomes allowed. The $^1\text{S}_0 \rightarrow ^1\text{P}_1$ is a spin allowed transition. Therefore, only the A- and C-bands have a high enough absorption strength to be used in phosphor applications. A more detailed discussion on the optical transitions of $6s^2$ ions can be found in the literature.^{3,4}

- Trends in bismuth luminescence as function of h-parameter. Applications: phosphors, scintillators, sensitizer for Eu (and other Ln) emission.

When Bi^{3+} is incorporated into a host lattice an additional absorption band is observed, which is often labeled as the D-band. This absorption originates from a metal-to-metal charge transfer (MMCT) transition, meaning that an electron from bismuth is transferred to the host cation $\text{Bi}^{3+}/\text{M}^{n+} \rightarrow \text{Bi}^{4+}/\text{M}^{(n-1)+}$. Recently, Boutinaud *et al.* developed a model to predict energy of the MMCT transition in d^0 and d^{10} transition-metal oxides doped with Bi^{3+} .^{5,6}

MMCT⁷

Location of energy levels determines optical properties and performance of devices. Com-

parison with lanthanides (Dorenbos model). 6s electrons are not shielded, unlike the 4f, therefore expected that the chemical environment has a critical influence on the location of the optical transitions of the bismuth ion.

In this paper we located the vacuum referred binding energies of the Bi^{3+} ion in a variety of host compounds.

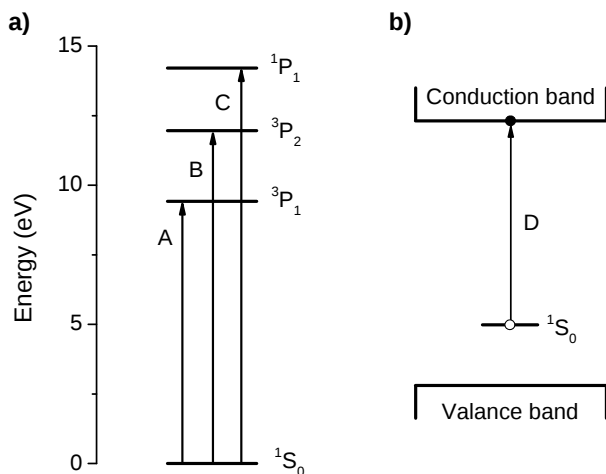


Figure 1: The energy levels of the free Bi^{3+} ion.

- Main focus paper: using MMCT to locate Bi^{3+} energy levels. Show that energy of s2 electron varies considerably with chemical environment.
- Bismuth self-quenching via pair emission as suggestion but focus for different paper. Potentially include bismuth as sensitizer.

Historical overview bismuth doped materials. What has been done: work of Blasse for phosphors, used in glass industry as probe ion,

Electronic configuration Bi^{3+} .

Bismuth pair formation. General for s2 elements in alkali halide crystals (and possibly all solids).

Incorporation into crystal results in an extra absorption/excitation band, the MMCT transition. Also depression of emission wavelength, very dependent on host lattice due to unshielded outer electrons.

Metal-to-metal charge transfer, useful for determining energy levels. Compare with IVCT of Pr^{3+} and Tb^{3+} .

Heavily-doped and self-activated bismuth compounds will be discussed in future work.

Dorenbos model: chemical shift, optical depression Ce^{3+} , VRBE

Redshift model, charge transfer model

Crystal field splitting and centroid shift only for Ce^{3+} or also applicable for Bi^{3+} ?

- Lanthanide free, which are expensive and only produced in China.
- How does s^2 luminescence work? Electron transitions, quantum mechanical splitting?
- Work of Blasse in the sixties.
- Dorenbos model on lanthanides.
- MMCT model Boutinaud.
- Paper of Wang, quantitative relation bismuth sp energy and host lattice.
- Bismuth as a sensitizer for Eu (and other Ln?) luminescence.
- Paper by Du: Chemical trends of electronic and optical properties of ns^2 ions in halides
- Optical electro negativity (Duffy)

Methodology

- How to locate bismuth energy levels? MMCT, A, C bands. B-band in most compounds too weak to be observed. Therefore excluded in this discussion.
- VRBE model
- Comparison with lanthanide spectroscopy: CT-bands, chemical shift model, redshift model, crystal field splitting, centroid shift.

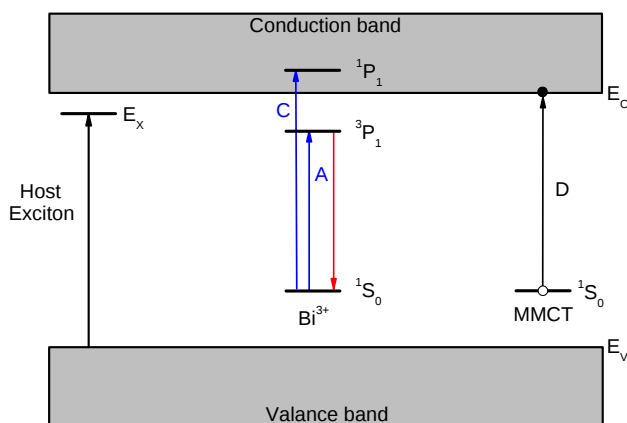


Figure 2: Electronic transitions in Bi^{3+} -doped compounds.

Results and discussion

Controleer toekenning A-band en D-band excitaties!

Aim of this paper: show Bi^{3+} energy levels in many compounds. Trends, how do these levels change with respect to chemical environment? Different emission bands, A-, C- and MMCT-band. Different excitations, A-band and MMCT. Difference between MMCT and C-band. In some compounds only MMCT, how to distinguish than between A-band?

- List (table) of all literature data found.
- VRBE schemes of Bi^{3+} in compounds.
- Compare with MMCT-model by Barandiaran *et al.*
- Pieter's model on energy level locations
- MMCT model Boutinaud *et al.*
- Example of concentration quenching (self-quenching/autoquenching)
- Comparison U-parameter and h-parameter: does it work for Bi^{3+} ?
- Stokes shift: calculate and show trends?
- Comparison with Eu and Ce. Eu at -4 eV and almost independent of host. Plot together

as function of U-parameter?

- Paper by Guifang: Y2O3 Bi3+ with lanthanides, how does sensitization work?
- Nephelauxetic sequence: increasing covalency, large effect on electron energies of 6s2 configuration, comparison with lanthanides.
- Shift of peak positions (A-band and MMCT excitations and A-band emission).
-

Data collected for 112 compounds. 7-digit compound identification number⁸

Data VRBE of host compounds from Dorenbos literature.

How does pair or mmct luminescence work? What is the emitting state?

Ju12 gave suggestion for mechanism bismuth sensitization of europium. We show that energy levels of excited state are at around -3.5 eV, matching with europium excited state!

Data on MMCT transition is rather scarce for wide band gap compounds because of limitation in excitation energy of most research group (200 nm limit).

Table 1: Spectroscopic data on Bi^{3+} in compounds.

ID number	Compound (A)	Excitation			Emission		Ref.
		A	C	D	A	CT	
0000000	free ion	75980	114610	—	—	—	9
1190010	BaF_2	46375	63880	75490	38320	—	10
1190020	SrF_2	46620	65815	73480	44120	—	10
1190030	CaF_2	47265	66140	73960	44440	—	10
1190404	NaYF_4	40330	50010	—	—	22745	11
2290002	RbCl	27780	43860	47620	—	—	12
2290003	KCl	30305	47160	49750	—	—	12
2290004	NaCl	30770	47170	50505	—	—	12
2290041	CsMgCl_3	35250	—	—	—	23875	13
2290101	$\text{Cs}_2\text{NaLaCl}_6$	31455	—	—	29280	—	14
2290401	$\text{Cs}_2\text{NaYCl}_6$	31000	—	—	30035	—	15
3390003	KBr	27030	42920	46295	—	—	12
3390071	CsCdBr_3	34925	—	—	17100	—	13
3390401	$\text{Cs}_2\text{NaYBr}_6$	27100	—	—	25490	—	14
4490003	KI	26315	41670	44845	—	—	12
5174020	$\text{Sr}_3\text{AlO}_4\text{F}$	32050	—	—	23420	—	16
5190400	YOF	37300	—	50000	30300	—	1
5290100	LaOCl	30000	—	37200	29000	22500	1
5290300	GdOCl	29600	—	38875	22500	20165	17
5290400	YOCl	30100	—	39200	25000	—	1
5390100	LaOBr	—	—	36535	27180	20165	18
5532100	LaP_3O_9	42500	—	—	—	21900	19
5532105	$\text{LiLaP}_4\text{O}_{12}$	43555	—	—	—	23310	20

Table 1: Continued

ID number	Compound (A)	Excitation			Emission		Ref.
		A	C	D	A	CT	
5532300	GdP ₃ O ₉	41500	—	—	—	—	19
5532400	YP ₃ O ₉	41500	—	—	34100	—	19
5532500	LuP ₃ O ₉	41500	—	—	34600	—	19
5532600	ScP ₃ O ₉	39500	—	—	35000	—	19
5534100	LaPO ₄	40815	—	57145	-	22220	21
5534400	YPO ₄	44445	—	58820	40985	29850	22
5534500	LuPO ₄	44445	—	—	42375	30030	22
5550013	KBaBP ₂ O ₈	41665	—	—	25840	—	23
5552100	LaB ₃ O ₆	38460	—	—	26315	—	15
5552140	LaMgB ₅ O ₁₀	33785	—	—	29760	—	17,24
5552160	LaZnB ₅ O ₁₀	33560	—	—	29760	—	24
5552170	LaCdB ₅ O ₁₀	33900	—	—	30300	—	24
5552300	GdB ₃ O ₆	—	—	—	—	—	15
5552440	YMgB ₅ O ₁₀	37315	—	—	30675	—	17,24
5552460	YZnB ₅ O ₁₀	37040	—	—	30120	—	24
5552470	YCdB ₅ O ₁₀	37315	—	—	30395	—	24
5554000	GaBO ₃	36215	—	—	34520	23630	25
5554000	InBO ₃	35210	—	—	32895	25000	26
5554035	LiCaBO ₃	32895	—	—	26455	—	27
5554100	LaBO ₃	37260	—	—	27910	21695	28
5554300	GdBO ₃	—	—	—	—	—	28
5554400	YAl ₃ B ₄ O ₁₂	38500	—	—	34500	—	1
5554400	YBO ₃	40485	—	54055	34015	31250	29

Table 1: Continued

ID number	Compound (A)	Excitation			Emission		Ref.
		A	C	D	A	CT	
5554500	LuBO ₃	34840	—	—	31850	—	26
5554600	ScBO ₃	34360	—	—	33390	—	28
5555430	CaYBO ₄	35800	—	—	—	—	30
5555430	Ca ₄ YO(BO ₃) ₃	32260	43480	—	26315	—	31
5563400	Y ₂ Sn ₂ O ₇	35715	—	—	30300	19610	32
5564025	Li ₄ SrCa(SiO ₄) ₂	32260	—	44445	32260	—	33
5564040	MgGeO ₃	34480	—	46510	27780	—	34
5564060	Zn ₂ GeO ₄	33330	—	—	—	19420	35
5564300	Gd ₂ GeO ₅	32260	—	42555	22220	—	36
5564405	LiYSiO ₄	35700	—	—	—	—	30
5565400	Y ₂ SiO ₅	36295	—	—	29035	16750	37
5565500	Lu ₂ SiO ₅	33875	40570	47990	27990	18150	38
5570000	ZnGa ₂ O ₄	27780	—	35715	24390	18520	39
5570100	LaAlO ₃	35090	—	—	26670	—	40
5570100	LaGaO ₃	32570	—	41670	26315	—	41
5570100	LaInO ₃	29400	—	—	23810	—	40
5570300	Gd ₃ Al ₅ O ₁₂	36100	—	—	26180	—	42
5570300	Gd ₃ Ga ₅ O ₁₂	34480	—	—	—	21280	43
5570400	Y ₃ Al ₅ O ₁₂	36495	—	49260	32950	21290	44
5570400	Y ₄ Al ₂ O ₉	33875	—	—	25810	—	45
5570400	YAlO ₃	35690	—	52500	29600	—	46
5570400	Y ₃ Ga ₅ O ₁₂	35200	—	—	31250	23810	47
5570500	Lu ₃ Al ₅ O ₁₂	36970	—	49875	33510	20650	44

Table 1: Continued

ID number	Compound (A)	Excitation			Emission		Ref.
		A	C	D	A	CT	
5573300	Gd ₂ GaSbO ₇	34480	—	—	27030	—	48
5581030	CaMoO ₄	—	—	30900	—	17540	5
5582030	CaWO ₄	—	—	34845	—	21370	49
5582060	ZnWO ₄	—	—	29410	—	17860	50
5582070	CdWO ₄	—	—	28570	—	18180	50
5582400	Y ₂ WO ₆	—	—	29300	—	19400	1
5583100	LaVO ₄	—	—	30860	—	18215	5
5583300	GdVO ₄	—	—	30490	—	17985	51
5583400	YVO ₄	—	—	30030	—	17545	51
5583500	LuVO ₄	—	—	29850	—	17360	52
5583600	ScVO ₄	—	—	28170	21505	15750	53
5584030	CaNb ₂ O ₆	—	—	31300	19610	—	5
5584100	LaNbO ₄	—	—	32790	24390	—	54
5584300	GdNbO ₄	—	—	32575	22470	—	55
5584400	YNbO ₄	—	—	31850	22520	18520	56
5585300	GdT _a ₇ O ₁₉	32260	—	—	20835	—	57
5585400	YT _a O ₄	34480	—	40000	23810	—	58
5586030	CaTiO ₃	—	—	27030	—	17240	59
5586400	Y ₂ Ti ₂ O ₇	—	—	31250	—	18180	22
5587030	CaZrO ₃	31250	—	—	25640	—	40
5587100	La ₂ Zr ₂ O ₇	34480	—	40820	25975	19420	60
5588030	CaHfO ₃	32500	—	—	26200	—	47
5589110	BaLa ₂ ZnO ₅	31250	—	37040	24390	—	61

Table 1: Continued

ID number	Compound (A)	Excitation			Emission		Ref.
		A	C	D	A	CT	
5590020	SrO	27260	—	37910	22990	—	62
5590030	CaO	28935	—	36455	25445	—	9
5590100	La ₂ O ₃	32470	—	40160	21980	20835	51
5590104	NaLaO ₂	28500	—	—	18000	—	63
5590300	Gd ₂ O ₃	28820	—	—	23530	18450	64
5590304	NaGdO ₂	29100	—	39500	26000	—	63
5590305	LiGdO ₂	30200	—	38900	21700	—	63
5590400	Y ₂ O ₃	30100	—	38000	24400	20800	1
5590404	NaYO ₂	28300	—	39400	26000	—	63
5590405	LiYO ₂	30800	—	—	18000	—	63
5590500	Lu ₂ O ₃	26860	—	30380	24800	19520	51
5590504	NaLuO ₂	28400	—	39100	26100	—	63
5590505	LiLuO ₂	30300	—	—	19500	—	63
5590600	Sc ₂ O ₃	29840	—	37100	24600	19840	65
5590604	NaScO ₂	27700	—	38700	26200	—	63
5590605	LiScO ₂	31600	—	39500	24700	—	63
6690020	SrS	23230	—	30245	20970	12500	62
6690030	CaS	24270	28795	32020	22220	16530	9
6690040	MgS	24035	28550	30890	22885	—	66
7790030	CaSe	22100	25970	29035	20200	15625	9

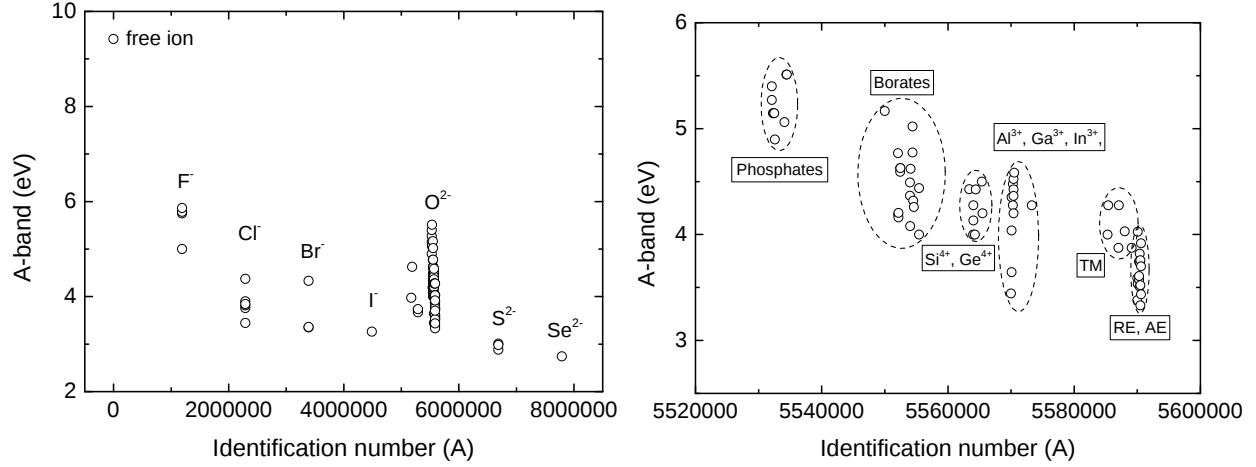


Figure 3: The A-band energies of Bi^{3+} in compounds.

Fig. 3 shows a decrease in the A-band absorption energy as function of the compound identification number (A). The strong decrease in the A-band transition energy from 9.4 eV in the free bismuth ion to 2.7 eV in selenide compounds is caused by a strong crystal field interaction of the 6s-electron with the chemical environment. The decrease follows the nephelauxetic sequence⁶⁷

$$\text{F}^- < \text{O}^{2-} < \text{Cl}^- < \text{Br}^- < \text{I}^- < \text{S}^{2-} < \text{Se}^- \quad (1)$$

$$\text{P}^{5+} < \text{B}^{3+} < \text{Si}^{4+} < \text{Al}^{3+} < \text{RE}^{3+} < \text{AE}^{2+} \quad (2)$$

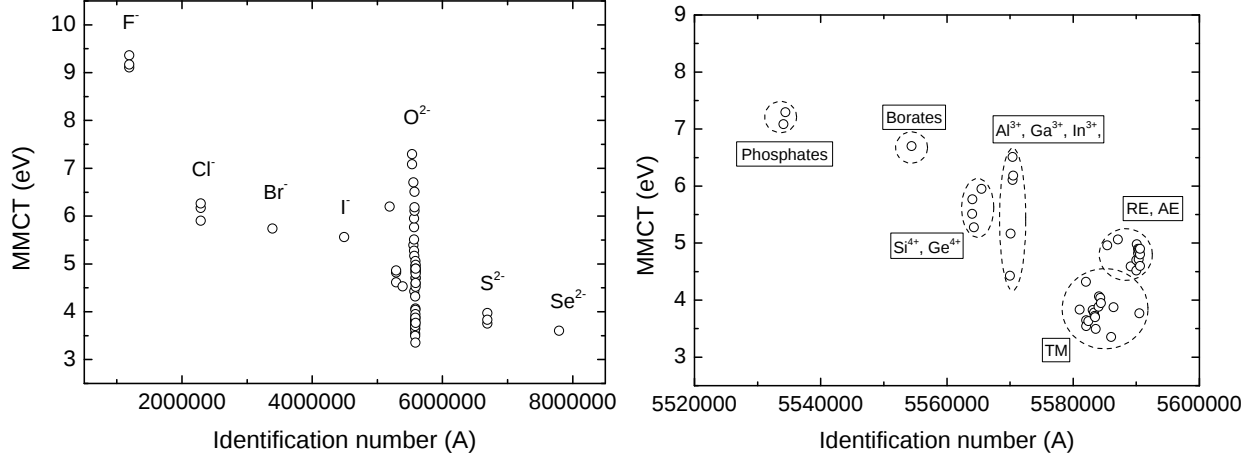


Figure 4: The MMCT energies of Bi^{3+} in compounds.

Eu^{3+} charge transfer energies in compounds⁶⁸

For Bi^{3+} doped in compounds containing transition metals (titanates, vanadates, niobates, tantalates, molybdates and tungstates) broad excitation and emission bands are observed. This is typical for charge transfer type of transition and was also observed by Boutinaud *et al.*⁵ These type of compounds have a low lying conduction band bottom and therefore in most of these compounds no interconfigurational transitions (A- or C-band) are observed, since the $^3\text{P}_1$ state is located inside or close to the conduction band bottom.

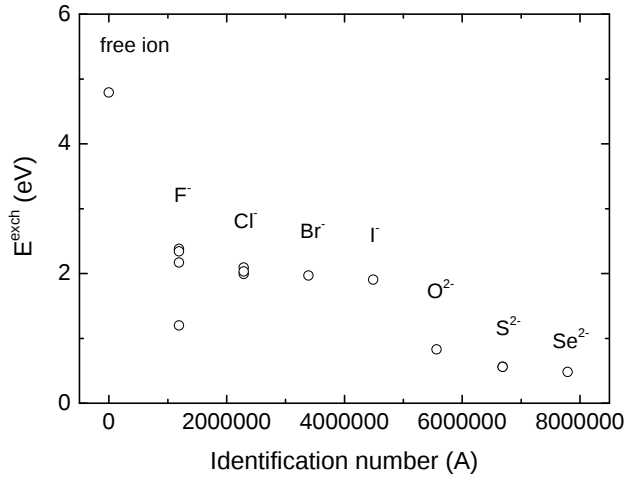


Figure 5: The exchange energies (E^{exch}) of Bi^{3+} in compounds.

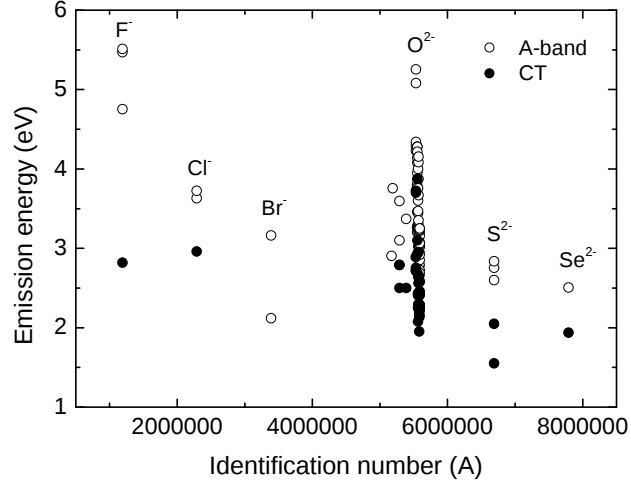


Figure 6: The A-band and CT emission energies in compounds.

The fact that the MMCT/pair emission is rather constant could indicate that the emission is not from cation- Bi^{3+} luminescence transition but from bismuth pairs (IVCT).

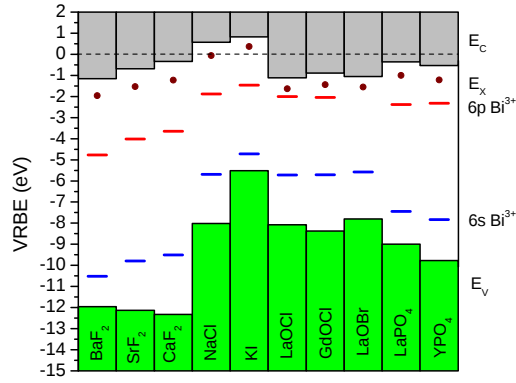


Figure 7: The vacuum referred binding energies of Bi^{3+} in compounds.

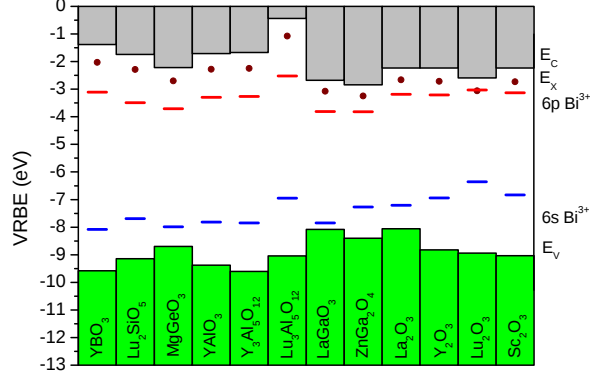


Figure 8: The vacuum referred binding energies of Bi^{3+} in compounds.

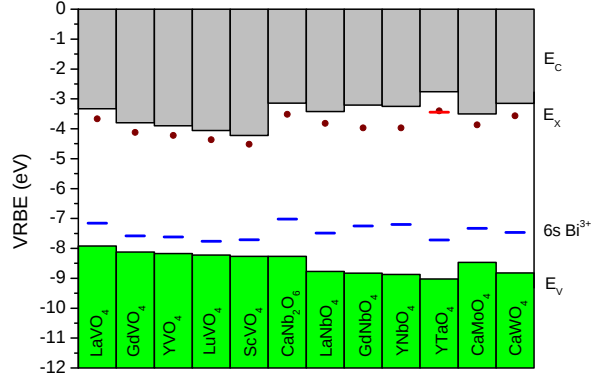


Figure 9: The vacuum referred binding energies of Bi^{3+} in compounds.

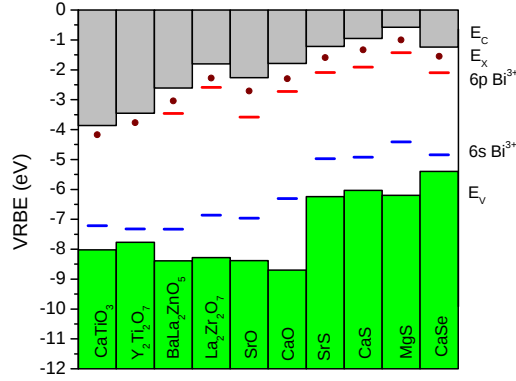


Figure 10: The vacuum referred binding energies of Bi^{3+} in compounds.

The chemical shift of the 6s-electron binding energy towards higher energy (less binding) as compared to the binding energy in the gaseous bismuth ion (free ion) is introduced by the crystal field of the host compound.

Conclusions

Acknowledgments

This work was supported by the Dutch Technology Foundation STW, which is part of the Netherlands Organisation for Scientific Research (NWO), and which is partly funded by the Ministry of Economic Affairs. This work was partly funded by Saint Gobain Crystals, France.

References

- (1) Blasse, G.; Bril, A. Investigations of Bi^{3+} -activated phosphors. *J. Chem. Phys.* **1968**, *48*, 217–222.
- (2) Srivastava, A. M.; Camardello, S. J. Concentration dependence of the Bi^{3+} luminescence in LnPO_4 ($\text{Ln} = \text{Y}^{3+}$, Lu^{3+}). *Opt. Mater.* **2015**, *39*, 130–133.
- (3) Ranfagni, A.; Mugnai, D.; Bacci, M.; Viliani, G.; Fontana, M. P. The optical properties of thallium-like impurities in alkali-halide crystals. *Adv. Phys.* **1983**, *32*, 823–905.
- (4) Jacobs, P. W. M. Alkali halide crystals containing impurity ions with the ns^2 ground-state electronic configuration. *J. Phys. Chem. Solids* **1991**, *52*, 35–67.
- (5) Boutinaud, P.; Cavalli, E. Predicting the metal-to-metal charge transfer in closed-shell transition metal oxides doped with Bi^{3+} or Pb^{2+} . *Chem. Phys. Lett.* **2011**, *503*, 239–243.
- (6) Boutinaud, P. Revisiting the spectroscopy of the Bi^{3+} ion in oxide compounds. *Inorg. Chem.* **2013**, *52*, 6028–6038.
- (7) Barandiarán, Z.; Meijerink, A.; Seijo, L. Configuration coordinate energy level diagrams

- of intervalence and metal-to-metal charge transfer states of dopant pairs in solids. *Phys. Chem. Chem. Phys.* **2015**, *17*, 19874–19884.
- (8) Dorenbos, P. The 5d level positions of the trivalent lanthanides in inorganic compounds. *J. Lumin.* **2000**, *91*, 155–176.
- (9) Yamashita, N.; Asano, S. Luminescence centers of $\text{Ca}(\text{S} : \text{Se}) : \text{Bi}^{3+}$ and $\text{CaO} : \text{Bi}^{3+}$ phosphors. *J. Phys. Soc. Japan* **1976**, *40*, 144–151.
- (10) Oboth, K. P.; Lohmeier, F. J.; Fischer, F. VUV and UV spectroscopy of Pb^{2+} and Bi^{3+} centres in alkaline-earth fluorides. *Phys. Stat. Sol. b* **1989**, *154*, 789–803.
- (11) Chong, K.; Hirai, T.; Kawai, T.; Hashimoto, S.; Ohno, N. Optical properties of Bi^{3+} ions doped in NaYF_4 . *J. Lumin.* **2007**, *122-123*, 149–151.
- (12) Radhakrishna, S.; Setty, R. S. S. Bismuth centers in alkali halides. *Phys. Rev. B* **1976**, *14*, 969–976.
- (13) Wolfert, A.; Blasse, G. Luminescence of s^2 ions in CsCdBr_3 and CsMg_3 . *J. Solid State Chem.* **1984**, *55*, 344–352.
- (14) Wolfert, A.; Blasse, G. Luminescence of Bi^{3+} -doped crystals of $\text{Cs}_2\text{NaYBr}_6$ and $\text{Cs}_2\text{NaLaCl}_6$. *J. Solid State Chem.* **1985**, *59*, 133–142.
- (15) van der Steen, A. C. Luminescence of $\text{Cs}_2\text{NaYCl}_6 - \text{Bi}^{3+}$ ($6s^2$). *Phys. Stat. Sol. b* **1980**, *100*, 603–611.
- (16) Noh, M.; Cho, S.-H.; Park, S. Tunable luminescence in Bi^{3+} and Eu^{3+} co-doped $\text{Sr}_3\text{AlO}_4\text{F}$ oxyfluorides phosphors. *J. Lumin.* **2015**, *161*, 343–346.
- (17) Wolfert, A.; Blasse, G. Luminescence of the Bi^{3+} ion in compounds LnOCl ($\text{Ln} = \text{La}, \text{Y}, \text{Gd}$). *Mater. Res. Bull.* **1984**, *19*, 67–75.

- (18) Wolfert, A.; Blasse, G. Luminescence of Bi^{3+} -activated LaOBr , a system with emission from different states. *J. Lumin* **1985**, *33*, 213–226.
- (19) Oomen, E. W. J. L.; Blasse, G. Luminescence of Bi^{3+} in the metaphosphates LnP_3O_9 ($\text{Ln} = \text{Sc, Lu, Y, Gd, La}$). *J. Solid State Chem.* **1988**, *75*, 201–204.
- (20) Babin, V.; Chernenko, K.; Demchenko, P.; Mihokova, E.; Nikl, M.; Pashuk, I.; Shalapska, T.; Voloshinovskii, A.; Zazubovich, S. Luminescence and excited state dynamics in Bi^{3+} -doped $\text{LiLaP}_4\text{O}_{12}$ phosphates. *J. Lumin.* **2016**, *176*, 324–330.
- (21) Moncorgé, R.; Boulon, G.; Denis, J. P. Fluorescence properties of bismuth-doped LaPO_4 . *J. Phys. C* **1979**, *12*, 1165–1171.
- (22) Srivastava, A. M.; Comanzo, H. A.; Camaradello, S. J. On the $\text{Bi}^{3+} - \text{Ti}^{4+}$ charge transfer transition in the pyrochlore $\text{Y}_2\text{Ti}_2\text{O}_7 : \text{Bi}^{3+}$. *Opt. Mater.* **2015**, *48*, 31–35.
- (23) Han, B.; Zhang, J.; Li, P.; Li, J.; Bian, Y.; Shi, H. Photoluminescence properties of novel $\text{KBaP}_2\text{O}_8:\text{M}$ ($\text{M} = \text{Pb}^{2+}$ and Bi^{3+}) phosphors. *Opt. Mater.* **2014**, *37*, 241–244.
- (24) Jagannathan, R.; Manoharan, S. P.; Rao, R. P.; Kutty, T. R. N. Luminescence and energy levels of Mn^{2+} in $\text{LnMB}_5\text{O}_{10}$ ($\text{Ln} = \text{La, Gd and Y}$; $\text{M} = \text{Mg, Zn and Cd}$). *Jpn. J. Appl. Phys.* **1990**, *29*, 1991–1996.
- (25) Dotsenko, V. P.; Efryushina, N. P.; Berezovskaya, I. B. Luminescence properties of $\text{GaBO}_3 : \text{Bi}^{3+}$. *Mater. Lett.* **1996**, *28*, 517–520.
- (26) Dotsenko, V. P.; Berezovskaya, I. B.; Efryushina, N. P. Photoionization and luminescence properties of Bi^{3+} in $\text{In}_{1-x}\text{Lu}_x\text{BO}_3$ solid solutions. *J. Phys. Chem. Solids* **1995**, *57*, 437–441.
- (27) Pekgözlü, I.; Erdoğan, E.; Çubuk, S.; Başak, A. S. Synthesis and photoluminescence of $\text{LiCaBO}_3 : \text{M}$ ($\text{M} : \text{Pb}^{2+}$ and Bi^{3+}) phosphor. *J. Lumin.* **2012**, *132*, 1394–1399.

- (28) Wolfert, A.; Oomen, E. W. J. L.; Blasse, G. Host lattice dependence of the Bi^{3+} luminescence in orthoborates LnBO_3 (with $\text{Ln} = \text{Sc}, \text{Y}, \text{La}, \text{Gd}, \text{or Lu}$). *J. Solid State Chem.* **1985**, *59*, 280–290.
- (29) Chen, L.; Zheng, H.; Cheng, J.; Song, P.; Yang, G.; Zhang, G.; Wu, C. Site-selective luminescence of Bi^{3+} in the YB_3 host under vacuum ultraviolet excitation at low temperature. *J. Lumin.* **2008**, *158*, 115–119.
- (30) Blasse, G. The ultraviolet absorption bands of Bi^{3+} and Eu^{3+} in oxides. *J. Solid State Chem.* **1972**, *4*, 52–54.
- (31) Ju, G.; Hu, Y.; Chen, L.; Wang, X.; Mu, Z.; Wu, H.; Kang, F. The luminescence of bismuth and europium in $\text{Ca}_4\text{YO}(\text{BO}_3)_3$. *J. Lumin.* **2012**, *132*, 717–721.
- (32) Srivastava, A. M. On the luminescence of Bi^{3+} in the pyrochlore $\text{Y}_2\text{Sn}_2\text{O}_7$. *Mater. Res. Bull.* **1999**, *37*, 745–751.
- (33) Pekgözlü, I.; Erdoğan, E.; Yilmaz, M. Synthesis and photoluminescence of $\text{Li}_4\text{SrCa}(\text{SiO}_4)_2: \text{M}$ ($\text{M}: \text{Pb}^{2+}$ and Bi^{3+}). *J. Lumin.* **2015**, *161*, 160–163.
- (34) Katayama, Y.; Ueda, J.; Tanabe, S. Effect of Bi_2O_3 doping on persistent luminescence of $\text{MgGeO}_3:\text{Mn}^{2+}$ phosphor. *Opt. Mater. Express* **2014**, *4*, 613–623.
- (35) Zhang, S.; Hu, Y.; Chen, R.; Wang, X.; Wang, Z. Photoluminescence and persistent luminescence in Bi^{3+} -doped Zn_2GeO_4 phosphors. *Opt. Mater.* **2014**, *36*, 1830–1835.
- (36) Guo, P.; Zhao, F.; Li, G.; Liao, F.; Tian, S.; Jing, X. Novel phosphors of Eu^{3+} , Tb^{3+} or Bi^{3+} activated Gd_2GeO_5 . *J. Lumin.* **2003**, *105*, 61–67.
- (37) Babin, V.; Gorbenko, V.; Krasnikov, A.; Mihokova, E.; Nikl, M.; Zazubovich, S.; Zorenko, Y. Photoluminescence and excited state structure in Bi^{3+} -doped Y_2SiO_5 single crystal films. *Radiat. Meas.* **2013**, *56*, 90–93.

- (38) Gorbenko, V.; Krasnikov, A.; Mihokova, E.; Nikl, M.; Zazubovich, S.; Zorenko, Y. Photoluminescence and excited state structure of Bi^{3+} -related centers in $\text{Lu}_2\text{SiO}_5\text{:Bi}$ single crystal films. *J. Lumin.* **2013**, *134*, 469–476.
- (39) Zhuang, Y.; Ueda, J.; Tanabe, S. Photochromism and white lon-lasting persistent luminescence in Bi^{3+} -doped ZnGa_2O_4 ceramics. *Opt. Express* **2012**, *2*, 1378–1383.
- (40) van Steensel, L. I.; Bokhove, S. G.; van de Craats, A. M.; de Blank, J.; Blasse, G. The luminescence of Bi^{3+} in LaInO_3 and some other perovskites. *Mater. Res. Bull.* **1995**, *30*, 1359–1362.
- (41) Jacquier, B.; Boulon, G.; Sallavuard, G.; Gaume-Mahn, F. Bi^{3+} center in a lanthanum gallate phosphor. *J. Solid State Chem.* **1972**, *4*, 374–378.
- (42) L.Tian,.; Wang, L.; Zhang, L.; Zhang, Q.; Ding, W.; Yu, M. Enhanced luminescence of $\text{Dy}^{3+}/\text{Bi}^{3+}$ co-doped $\text{Gd}_3\text{Al}_5\text{O}_{12}$ phosphors by high-efficiency energy transfer. *J. Mater. Sci: Mater. Electron* **2015**, *26*, 8507–8514.
- (43) Novoselov, A.; Yoshikawa, A.; Nikl, M.; Solovieva, N.; Fukuda, T. Shaped single crystal growth and scintillation properties of $\text{Bi:Gd}_3\text{Ga}_5\text{O}_{12}$. *Nucl. Instrum. Meth. Phys. Res. A* **2005**, *537*, 247–250.
- (44) Zorenko, Y.; Mares, J. A.; Kucerkova, R.; Gorbenko, V.; Savchyn, V.; Voznyak, T.; Nikl, M.; Beitlerove, A.; Jurek, K. Optical, luminescence and scintillation characteristics of Bi-doped LuAG and YAG single crystalline films. *J. Phys. D: Appl. Phys.* **2009**, *42*, 075501.
- (45) Babin, V.; Lipińska, L.; Mihokova, E.; Nikl, M.; Shalapska, T.; Suchocki, A.; Zazubovich, S.; Zhydachevskii, Y. **2015**,
- (46) Krasnikov, A.; Lipińska, L.; Mihokova, E.; Nikl, M.; Shalapska, T.; Suchocki, A.; Za-

- zubovich, S.; Zhydachevskii, Y. Time-resolved photoluminescence and excited state structure of Bi^{3+} center in YAlO_3 . *Opt. Mater.* **2014**, *36*, 1705–1708.
- (47) Setlur, A. A.; Srivastava, A. M. The nature of Bi^{3+} luminescence in garnet hosts. *Opt. Mater.* **2006**, *29*, 410–415.
- (48) Srivastava, A. M.; Szarowski, A. On the quenching of Bi^{3+} luminescence in the pyrochlore $\text{Gd}_2\text{GaSbO}_7$. *J. Solid State Chem.* **1999**, *146*, 494–498.
- (49) Zorenko, Y.; Pashkovsky, M.; Voloshinovskii, A.; Kuklinski, B.; Grinberg, M. The luminescence of CaWO_4 single crystals. *J. Lumin* **2006**, *116*, 43–51.
- (50) Wang, L.; Lv, Z.; Kang, W.; Shangguan, X.; Shi, J.; Hao, Z. Applications oriented design of Bi^{3+} doped phosphors. *Appl. Phys. Lett.* **2013**, *102*, 151909.
- (51) Boulon, G. Processus de photoluminescence dans les oxydes et les orthovanadates de terres rares polycristallins actives par l'ion Bi^{3+} . *J. Phys. (Paris)* **1971**, *32*, 333–347.
- (52) Kang, F.; Peng, M.; Zhang, Q.; Qiu, J. Abnormal anti-quenching and controllable multi-transitions of Bi^{3+} luminescence by temperature in a yellow-emitting $\text{LuVO}_4\text{:Bi}^{3+}$ phosphor for UV-converted white LEDs. *Chem. Eur. J.* **2014**, *20*, 11522–11530.
- (53) Kang, F.; Yang, X.; Peng, M.; Wondraczek, L.; Ma, Z.; Zhang, Q.; Qiu, J. Red photoluminescence from Bi^{3+} and the influence of the oxygen-vacancy perturbation in ScVO_4 : A combined experimental and theoretical study. *J. Phys. Chem. C* **2014**, *118*, 7515–7522.
- (54) Park, T. K.; Ahn, H. C.; Mho, S. I. High concentration of Bi^{3+} incorporated into $\text{RNbO}_4\text{:Eu}^{3+}$ ($\text{R} = \text{La}, \text{Y}, \text{Gd}$) as red phosphors for white LED applications. *J. Korean Phys. Soc.* **2008**, *52*, 431–434.
- (55) Liu, X. M.; Lin, J. Enhanced luminescence of gadolinium niobates by Bi^{3+} doping for field effect emission displays. *J. Lumin.* **2007**, *122-123*, 700–703.

- (56) Shin, S. H.; Jeon, D. Y.; Suh, K. S. Charge-transfer nature in luminescence of $\text{YNbO}_4\text{:Bi}$ blue phosphor. *J. Appl. Phys.* **2001**, *90*, 5986–5990.
- (57) Kubota, S. I.; Yamane, H.; Shimada, M. Luminescence properties of $\text{Gd}_{1-x}\text{Bi}_x\text{Ta}_7\text{O}_{19}$. *J. Alloys Compds.* **1998**, *281*, 181–185.
- (58) Blasse, G.; Bril, A. Luminescence phenomena in compounds with fergusonite structure. *J. Lumin.* **1970**, *3*, 109–131.
- (59) Boutinaud, P.; Cavalli, E.; Mahiou, R. Photon conversion in $\text{Bi}^{3+}/\text{Pr}^{3+}$ -codoped CaTiO_3 . *J. Phys.: Condens. Matter* **2012**, *24*, 295502.
- (60) Srivastava, A. M.; Beers, W. W. On the impurity trapped exciton luminescence in $\text{La}_2\text{Zr}_2\text{O}_7\text{:Bi}^{3+}$. *J. Lumin* **1999**, *81*, 293–300.
- (61) Chang, Y.-S. Blue Emitting Phosphors of $\text{BaLa}_2\text{ZnO}_5$ Activated by Bismuth ions. *J. Electrochem. Soc.* **2011**, *128*, 2027–2030.
- (62) Ellervee, A. F. Luminescence of Pb^{2+} and Bi^{3+} centres in alkali-earth sulphides and oxides. *Phys. Stat. Sol. b* **1977**, *82*, 91–98.
- (63) van der Steen, A. C.; van Hesteren, J. J. A.; Slok, A. P. Luminescence of Bi^{3+} ion in LiLnO_2 and NaLnO_2 ($\text{Ln} = \text{Sc, Y, La, Gd, Lu}$). *J. Electrochem. Soc.* **1981**, *128*, 1327–1333.
- (64) Liu, G. X.; Zhang, R.; Xiao, Q. L.; Zou, S. Y.; Peng, W. F.; Cao, L. W.; Meng, J. X. Efficient $\text{Bi}^{3+} \rightarrow \text{Nd}^{3+}$ energy transfer in $\text{Gd}_2\text{O}_3\text{:Bi}^{3+}, \text{Nd}^{3+}$. *Opt. Mater.* **2011**, *34*, 313–316.
- (65) Bordun, O. M. Luminescence of bismuth-activated ceramics of yttrium and scandium oxides. *J. Appl. Spectrosc.* **2002**, *69*, 67–71.
- (66) Asano, S.; Yamashita, N. Luminescence et interaction phonon-electron dans le luminophore MgS:Bi^{3+} . *Phys. Stat. Sol. b* **1981**, *105*, 305–310.

- (67) Dorenbos, P. Lanthanide 4f-electron binding energies and the nephelauxetic effect in wide band gap compounds. *J. Lumin.* **2013**, *136*, 122–129.
- (68) Dorenbos, P. The Eu^{3+} charge transfer energy and the relation with the band gap of compounds. *J. Lumin.* **2005**, *111*, 89–104.

Graphical TOC Entry

