Insight into Eu redox and Pr$^{3+}$ 5d emission in KSrPO$_4$ by VRBE scheme construction†

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Series of Ln-doped KSrPO$_4$ (Ln = Ce$^{3+}$, Eu$^{2+}$, Eu$^{3+}$, Pr$^{3+}$) phosphors are prepared through a high-temperature solid-state method. The KSrPO$_4$ compound is confirmed to be $\beta$-K$_2$SO$_4$ structure with $Pmn$ group by Rietveld refinement, and the temperature-dependent lattice parameters are investigated with the powder X-ray diffraction results at different temperatures. Ce$^{3+}$ and Eu$^{3+}$ ions are introduced to probe the crystal field strength (CFS) and the lanthanide site symmetry by VUV-UV-vis spectroscopy. The temperature-dependent luminescence properties of KSrPO$_4$: Ce$^{3+}$/Eu$^{3+}$ exhibit excellent thermal stability of Ce$^{3+}$/Eu$^{3+}$ luminescence. Based on the VUV-UV-vis spectra of Ce$^{3+}$ and Eu$^{3+}$ doped KSrPO$_4$, the vacuum referred binding energy (VRBE) scheme is constructed to understand the redox properties of Eu, the 5d energy levels of Pr$^{3+}$ and thermal quenching characteristics of Ce$^{3+}$ and Eu$^{3+}$ luminescence.

1. Introduction

Eu$^{2+}$-activated luminescent materials have drawn wide attention because of their importance in solid-state lighting and displays.$^{1-4}$ The preparation of divalent europium doped materials often needs reducing conditions, for example, H$_2$ or CO ambient, etc. But in some special cases, Eu$^{3+}$ can also be (partially) obtained through a solid-state reaction approach with Eu(III)-containing raw materials in non-reducing conditions.$^5$ This is, of course, the result of composition and structure of these special host compounds. From the viewpoint of electron transfer, the redox processes of Eu$^{2+}$ and Eu$^{3+}$ as dopants in solid state inorganic compounds relate to factors such as the electronic structure of conduction band and valence band, the band gap energy and the Fermi level of the host compound, the 5d energy of Eu$^{2+}$, and the energy of charge transfer state (CTS) between Eu$^{3+}$ and coordinated atoms.$^7$ Clear understanding of the redox behaviour of Eu$^{2+}$ / Eu$^{3+}$ in different host compounds is a key issue for the preparation and application of Eu$^{2+}$-activated luminescent materials.

Eu$^{2+}$, Pr$^{3+}$ and Ce$^{3+}$ doped luminescent materials are also applied as important scintillators. SrI$_2$: Eu$^{2+}$, and LnX$_2$: Ce$^{3+}$ (Ln = La, Lu; X = Br, I) are commercially available scintillators in various detecting areas.$^8$-10 These applications are benefited from the typical parity-allowed 5d-4f transitions of these ions and therefore with fast radiative rates. In the same site of the same host compound, fluorescence decay of Pr$^{3+}$ is usually faster than that of Ce$^{3+}$ or Eu$^{3+}$ due to its shorter emission wavelength.$^{11}$ Therefore Pr$^{3+}$ is an important activator in fast decay scintillators. Because the binding energy of outer 5d electrons is highly sensitive to the local coordination environment of the host lattice, it is essential to fully understand the energy and thermal stability of d-f emission of Pr$^{3+}$ in various host compounds to find novel Pr$^{3+}$ doped scintillation materials.

The compound KSrPO$_4$ is one of the ABPO$_4$ (A = monovalent cations, B = divalent cations) monophosphates with tetrahedral rigid anion groups.$^{12}$ And Eu$^{2+}$, Tb$^{3+}$, Sm$^{2+}$/ Sm$^{3+}$ doped KSrPO$_4$ phosphors exhibit excellent luminescence properties.$^{13}$ In this paper, we investigate the temperature- and doping concentration-dependent synchrotron radiation VUV-UV-vis luminescence spectra of Ce$^{3+}$, Pr$^{3+}$, Eu$^{3+}$ and Eu$^{3+}$/O$^{2-}$ charge transfer energy with the purpose to understand the excellent thermal stabilities of luminescence of Ce$^{3+}$ and Eu$^{3+}$, the 5d energy of Pr$^{3+}$ and the redox of Eu. Eu-doped samples are prepared in inert N$_2$, oxidizing air and reducing CO conditions to get clearer insight into the redox characteristics of Eu in KSrPO$_4$. The work demonstrates a convenient approach to predict the stability of dopant Eu$^{2+}$ / Eu$^{3+}$ and the energy of d-f emission of Pr$^{3+}$ in a specific host compound.
2. Experimental

Ln-doped KSrPO$_4$ (Ln = Ce$^{3+}$, Eu$^{3+}$, Eu$^{2+}$, Pr$^{3+}$) samples are prepared using a conventional solid-state reaction route. The analytic reagents SrCO$_3$, K$_2$CO$_3$, NH$_4$H$_2$PO$_4$ and 99.99% pure rare-earth oxides Eu$_2$O$_3$, CeO$_2$, Pr$_2$O$_3$ are used as starting materials. These reactants are thoroughly ground in an agate mortar and pre-fired at 975 K in air atmosphere for 5 h. Then Ce$^{3+}$ or Pr$^{3+}$ doped samples are annealed at 1570 K for 3 h under CO ambiance which are produced from the incomplete combustion of carbon at high-temperature. Europium doped sample batches with each a specific concentration are divided into three parts and fired in reducing CO, inert N$_2$, or oxidizing air, respectively. Finally, the samples are cooled to RT by switching off the muffle furnace and ground into powders for subsequent analysis.

The phase purity of the samples is estimated by powder X-ray diffraction (P-XRD) on a Bruker D8 advanced X-ray diffractometer with a wavelength of 1.54056 Å Cu Kα radiation at 40 kV and 40 mA, demonstrating no detectable impurity phase in all samples. High quality P-XRD data for the refinement are collected within the 2θ range from 7° to 110° at a 2θ step of 0.02°. The Rietveld refinement is performed using the TOPAS - Academic program.$^{15}$

The UV-vis excitation and emission spectra as well as the luminescence decay curves are recorded at an Edinburgh FLS 920 combined fluorescence lifetime and steady-state spectrometer equipped with a thermo-electronic cooled (-19.5 °C) Hamamatsu R928P Model photomultiplier. The 450 W xenon lamp is used as the excitation source of steady-state excitation and corresponding emission spectra. A 150 W nF900 lamp with a pulse width of 1 ns and a 60 W μF flash lamp with a pulse width of 1.5-3.0 μs are used for the measurements of decay curves. The temperature dependent spectral measurements at 77-500 K range are performed by putting the samples in a liquid nitrogen cooling OptistatDN2 cryostat and controlled by an Oxford MercuryITC temperature controller. The excitation and emission spectra in the VUV range are measured on the beam line 4B8 of the Beijing Synchrotron Radiation Facility (BSRF) under normal operating conditions using the spectrum of sodium salicylate (o-C$_6$H$_5$OHCOONa) as a standard.$^{15,17}$

3. Results and discussion

3.1. Structure and Phase Characterization

![Fig. 1 Experimental (red crosses) and calculated (green solid line) XRD patterns, their difference (cyan solid line), and the Bragg reflection positions (blue ticks) of KSrPO$_4$, the inset shows the crystal structure of KSrPO$_4$.](image)

Fig. 1 shows Rietveld refinement results based on laboratory P-XRD data of the synthesized KSrPO$_4$ compound using the $Pnma$ structure model.$^{18,19}$ No impurity peak is observed, and the obtained reliability factors $R_{wp}$, $R_p$ and $R_B$ are 5.89%, 3.57%, and 5.92%, respectively. The final refined structural parameters are listed in ESI Table S1†. The KSrPO$_4$ compound is crystallized in an orthorhombic structure with space group $Pnma$. The lattice parameters are $a = 7.34706(4)$ Å, $b = 5.55249(3)$ Å, $c = 9.61716(6)$ Å and $V = 392.325(4)$ Å$^3$. The inset of Fig. 1 depicts the crystal structure of KSrPO$_4$. A Sr$^{2+}$ atom is coordinated by nine oxygen atoms with $C_c$ point symmetry. The average bond length between Sr$^{2+}$ and O$^2-$ is ~2.688 Å (ESI Table S2†) and the nearest distance of two adjacent Sr$^{2+}$ ions is ~3.809 Å. The polyhedral SrO$_9$, KO$_{10}$ and PO$_4$ form the framework of KSrPO$_4$ structure by sharing their corner or edges. The ESI Fig. S1† shows the representative P-XRD patterns of Ce$^{3+}$, Eu$^{3+}$, Pr$^{3+}$ and Eu$^{2+}$ singly doped KSrPO$_4$ samples. All XRD patterns are in good agreement with the refined KSrPO$_4$ result, implying that all samples are of single phase and the doping of Ce$^{3+}$/ Eu$^{3+}$/ Pr$^{3+}$/ Eu$^{2+}$ ions do not significantly influence the diffractogram of the host compound. Due to the more suitable ionic size and valence state \(r(\text{Ce}^{3+}) = 1.336 \text{ Å}, r(\text{Eu}^{3+}) = 1.26 \text{ Å}, r(\text{Pr}^{3+}) = 1.319 \text{ Å}, r(\text{Eu}^{2+}) = 1.44 \text{ Å}, r(\text{Sr}^{2+}) = 1.45 \text{ Å}\), coordination number = 9$^{20}$, Ce$^{3+}$, Eu$^{3+}$, Pr$^{3+}$ and Eu$^{2+}$ may prefer occupying the Sr$^{2+}$ site when doped into KSrPO$_4$. 
To study the thermal evolution of the host structure, Fig. 2 (a) displays P-XRD patterns of KSrPO$_4$ at temperatures from 223 to 523 K. KSrPO$_4$ maintains the orthorhombic structure (space group Pnma) in the investigated temperature range. The magnified patterns from 31.9 to 32.2 degree of 2-theta in Fig. 2 (b) shows that the diffraction peaks slightly shift to the lower angle side, which implies that the unit cell of the host undergoes a thermal expansion. Based on these XRD patterns, the lattice constants can be calculated by Rietveld refinement. As displayed in Fig. 2 (c), the lattice parameters $a$, $b$, $c$ and $V$ are found to increase with rising temperature, which may affect the luminescence properties of samples as discussed later.

3.2. Luminescence of Ce$^{3+}$ in KSrPO$_4$

Fig. 2 (a) Temperature-dependent P-XRD patterns of KSrPO$_4$ in the temperature range of 223-523 K; (b) magnified P-XRD patterns (31.9-32.2 degree of 2-theta) of KSrPO$_4$; (c) lattice parameters $a$, $b$, $c$ and $V$ of KSrPO$_4$ as a function of temperature.

Fig. 3 shows the VUV-UV excitation ($\lambda_{\text{exc}} = 332$ nm) spectrum of the sample with nominal composition K$_{1.001}$Sr$_{0.998}$Ce$_{0.001}$PO$_4$ at 25 K. The high-lying excitation band with a maximum at about 7.65 eV (band H) is attributed to the excitonic absorption of the KSrPO$_4$ host. The absorptions below ~6.50 eV are assigned to the f-d excitation bands of Ce$^{3+}$ in KSrPO$_4$. In this range three clear bands A, C, and D can be observed. The bands A and D have shoulder bands B and E at their higher energy sides, respectively. When we assume that Ce$^{3+}$ ions occupy nine-fold coordinated Sr$^{2+}$ sites with a low symmetry $C_j$ group, the 5d orbitals of Ce$^{3+}$ should be split into five non-degenerate orbitals. Therefore, the excitation curve in the 3.87-6.55 eV (190-320 nm) spectral range is fitted by a sum of five Gaussian functions to estimate the energies of 4f-5d transitions of Ce$^{3+}$. These energies are at about 4.14 (A), 4.60 (B), 5.23 (C), 5.76 (D), and 6.08 (E) eV, respectively. Consequently, the centroid energy of Ce$^{3+}$ 5d states, viz. the arithmetic mean energy of the five f-d excitation bands, is calculated to be about 5.16 eV, meaning that the down-shift of centroid energy of Ce$^{3+}$ 5d state ($\epsilon_f$) in KSrPO$_4$ is about 1.19 eV with respect to that of free gaseous Ce$^{3+}$ (6.35 eV). This value is near those of other phosphates, such as K$_2$LuPO$_4$ (1.17 eV),$^{21}$ YPO$_4$ (1.19 eV),$^{21}$ Ca$_5$Y(PO$_4$)$_2$ (1.20 eV),$^{22}$ and LuPO$_4$ (1.20 eV),$^{21}$ which implies that Ce$^{3+}$ ions in these phosphate compounds have nearly same nephelauxetic effect, covalence or spectroscopic polarization.$^{21,23}$ Meanwhile, the experimental crystal field splitting of Ce$^{3+}$ in KSrPO$_4$ is calculated to be approximately 1.94 eV by subtracting the energy of the first 5d orbital (4.14 eV) from that of the fifth 5d one (6.08 eV). This value is in the range of those of Ce$^{3+}$ in compounds SrB$_2$O$_7$ (1.66 eV),$^{24}$ Sr$_2$Mg(BO$_3$)$_2$ (2.10 eV),$^{25}$ and SrAl$_2$O$_4$ (2.68 eV, 2.75 eV),$^{26}$ in which Ce$^{3+}$ ions occupy nine-fold Sr$^{2+}$ sites with $C_j$ or $C_{ij}$ symmetry. The emission ($\lambda_{\text{em}} = 300$ nm) spectrum of K$_{1.001}$Sr$_{0.998}$Ce$_{0.001}$PO$_4$ at 25 K is fitted using a sum of two Gaussian components as displayed in the inset of Fig. 3. The obtained two bands (1, ~3.49 eV; 2, ~3.74 eV) are originated from the transitions from the lowest 5d state to its $^2$F$_j$ ($j = 7/2, 5/2$) 4f ground state of Ce$^{3+}$, respectively. Their energy difference is evaluated about 0.25 eV (2.02x$10^3$ cm$^{-1}$), which is coincident with the theoretical energy difference of $^2$F$_j$ ($j = 5/2, 7/2$) multiplets (~2.00-10$^3$ cm$^{-1}$). According to the energy difference between the band peaks of the lowest 5d and $^2$F$_{5/2}$ transition in excitation and emission spectra, the Stokes shift of Ce$^{3+}$ is approximately 0.40 eV (~3.23x$10^3$ cm$^{-1}$). Stokes shift is related to Huang-Rhys parameter and effective phonon energy and depends on the characteristic of material. The value in present case is close to that in other monophosphates, such as NaCaPO$_4$ (0.38 eV), and Li$_5$SrPO$_4$ (0.40 eV).$^{27,28}$ The ESI Fig. S2† shows the highest-height normalized VUV-UV excitation ($\lambda_{\text{exc}} = 332$ and 354 nm) and emission ($\lambda_{\text{em}} = 200, 215, 235, 270$ and 300 nm) spectra of sample K$_{1.001}$Sr$_{0.998}$Ce$_{0.001}$PO$_4$ at 25 K. The excitation spectra by monitoring different emission wavelengths overlap to each other, and the emission spectra under different wavelength excitation also have the similar profiles. These phenomena demonstrate that the sample is without impurity phase and reveal the occurrence of only one
kind of Ce$^{3+}$ luminescence center in the sample as we mentioned above. The luminescence decay curves at different wavelength conditions in ESI Fig. S3† further confirm this result. They all possess the same exponential decay properties and the lifetime of Ce$^{3+}$ in KSrPO$_4$ is evaluated about 25.7 ns.

The height-normalized excitation ($\lambda_{em} = 375$ nm) and emission ($\lambda_{ex} = 270$ nm) spectra of samples K$_{1-x}$Sr$_{1.2x}$Ce$_x$PO$_4$ ($x = 0.001, 0.003, 0.009, 0.019$ and $0.03$) are plotted in Fig. 4. With increasing Ce$^{3+}$ content, excitation band of $4f$-$5d_1$ transition and emission spectra gradually shift to lower energies, and the intensity of $5d_1^{2}F_{5/2}$ emission increases with respect to that of $5d_1^{2}F_{7/2}$. Since smaller Ce$^{3+}$ ion occupy larger Sr$^{2+}$ site, the lattice may experience a little contraction and the CFS around Ce$^{3+}$ becomes larger, which pushes the $5d_1$ level towards lower energies. The zero phonon line (ZPL) position directly indicates the energy of the $4f$-$5d_1$ transition of Ce$^{3+}$. Its position can be estimated as the intersection point of the excitation and emission spectra, although we did not observe ZPLs in the spectra. Fig. 4 shows the red shift of the intersection point with increasing Ce$^{3+}$ doping, indicating that ZPLs shift to lower energies with stronger CFS around Ce$^{3+}$. Accordingly, the emission peaks of Ce$^{3+}$ gradually shift to longer wavelengths. Moreover, in consideration of the significant overlap between excitation and emission spectra, the self-absorption effect of Ce$^{3+}$ is also active in our case. This effect also shifts the emission band to the lower energy side in some extent, and meanwhile gives rise to the decrease of relative intensity of $5d_1^{2}F_{5/2}$ emission. Therefore, the relative intensity of $5d_1^{2}F_{7/2}$ emission looks stronger in height-normalized spectra. The ESI Fig. S4† (a) shows the concentration-dependent emission spectra of K$_{1-x}$Sr$_{1.2x}$Ce$_x$PO$_4$ ($x = 0.001, 0.003, 0.009, 0.019$ and $0.03$) samples under 270 nm excitation at RT. With increase of doping contents, emission intensities of Ce$^{3+}$ increase gradually, demonstrating that concentration quenching does not occur in this concentration range. The luminescence decay curves ($\lambda_{ex} = 290$ nm, $\lambda_{em} = 354$ nm) of samples are shown in ESI Fig. S4† (b). All curves follow the exponential characteristic and overlapped with each other, further confirming that concentration quenching of Ce$^{3+}$ emission does not occur until $x = 0.03$.

Fig. 5, displays the normalized emission ($\lambda_{ex} = 290$ nm) spectra and the decay curves ($\lambda_{ex} = 290$ nm, $\lambda_{em} = 354$ nm) of sample K$_{1.000}$Sr$_{0.999}$Ce$_{0.001}$PO$_4$ at different temperatures. From 78 to 500 K, the emission band shows a bit of short-wavelength shifting, because of the lattice expansion resulted weaker crystal field strength at higher temperatures. Moreover, the intensities of Ce$^{3+}$ $5d_1^{2}F_{1/2}$ ($J = 5/2, 7/2$) transitions decrease gradually, and the high-energy $5d_1^{2}F_{7/2}$ emission reduces relatively faster than the low-energy $5d_1^{2}F_{5/2}$ as seen in Fig. 5 (a). This is due to the enhanced reabsorption at higher temperatures as illustrated in ESI Fig. S5†. Meanwhile, two partial-resolved emissions become more overlapped because of the thermal-broadening of both bands. All decay curves in Fig. 5 (b) show exponential properties and no shortening lifetime with rising temperature, implying that the thermal quenching of Ce$^{3+}$ emission does not occur yet. Instead, the slight increase of Ce$^{3+}$ decay time can be observed in the inset of Fig. 5 (b). In consideration of the self-absorption at higher temperature, the lifetime of Ce$^{3+}$ can be slightly lengthened.

3.3. Luminescence of Eu$^{3+}$ in KSrPO$_4$ and Construction of the VRBE Scheme
Fig. 6 displays the synchrotron radiation VUV-UV excitation spectrum \((\lambda_{ex} = 610 \text{ nm})\) and laboratory UV-vis excitation spectrum \((\lambda_{ex} = 611 \text{ nm})\) of \(K_{1.003}Sr_{0.99}Eu_{0.007}PO_4\) sample at RT. The peak position and band shape of the high energy excitation band \((162 \text{ nm}, 7.65 \text{ eV})\) is the same with that in Fig. 3, and this band definitely corresponds to the excitonic absorption of this phosphate host. For the electron and hole binding energy in the exciton we will assume a value of 0.008\((e^\text{H})^3\) as proposed in reference.\(^{31}\) Accordingly, the bottom of the conduction band is estimated about 7.65 + 0.47 = 8.12 eV higher than the top of the valence band. The intense broad excitation band at about 272 nm \((4.56 \text{ eV})\) is assigned to the Eu\(^{3+}\)-O\(^2-\) charge transfer transition. A series of sharp lines due to the f-f transition of Eu\(^{3+}\) can be observed below 300 nm.

Fig. 7 (a) shows the emission spectrum \((\lambda_{es} = 393 \text{ nm})\) of \(K_{1.003}Sr_{0.99}Eu_{0.007}PO_4\) sample at RT, which contains a series of 4f intraconfigurational transitions from the excited \(D_2\) level to \(F_j\) \((j = 0, 1, 2, 3, 4 \text{ and } 5)\) and the strongest emission is \(D_0\)-\(F_2\) transition with a maximum at 609.3 nm. The emission intensity of \(D_0\)-\(F_2\) transition is hypersensitive to the Eu\(^{3+}\) microenvironment and depends closely on the point symmetry of Eu\(^{3+}\) site. The occurrence of the dominant \(5D_0\)-\(F_2\) transition indicates that Eu\(^{3+}\) ion is incorporated into a site without inversion symmetry. The main \(5D_0\)-\(F_2\) transition together with strong \(5D_1\)-\(F_4\) transition results in the deep-red luminescence of \(K_{1.003}Sr_{0.99}Eu_{0.007}PO_4\) phosphor with CIE chromaticity coordinates \((0.634, 0.365)\) as exhibited in Fig. 7 (b). Meanwhile, the lifetime is measured to be 2.5 ms as shown in the inset of Fig. 7 (a).

Provided that Eu\(^{3+}\) ions occupy Sr\(^{2+}\) sites with \(C_3\) point symmetry, the degeneracy of each \(F_{0.1.2}\) multiplet of Eu\(^{3+}\) would be completely lifted. To verify the site symmetry of Eu\(^{3+}\) in KSrPO\(_4\), in Fig. 7 (c) we give a more detailed emission spectrum of sample \(K_{1.003}Sr_{0.99}Eu_{0.007}PO_4\) in the 15750-17500 cm\(^{-1}\) range, which is measured with step size 0.05 nm under the 531 nm \(F_0\)-\(D_1\) transition excitation at 3 K. Since the \(5D_0\)-\(F_2\) transition will not be split, it can be used to determine the number of sites occupied. The single weak sharp line at 17350 cm\(^{-1}\) of the \(5D_0\)-\(F_2\) transition confirms one specific Eu\(^{3+}\) site occupation in KSrPO\(_4\). Three emission lines in 16715-17160 cm\(^{-1}\) range are the results of magnetic-dipole \(5D_0\)-\(F_2\) transitions, and the lines in the range from 15850 to 16550 cm\(^{-1}\) are ascribed to the electric-dipole \(5D_0\)-\(F_2\) transitions. The \(5D_0\)-\(F_2\) emission splits into four clear lines and one line trailing in the low wavenumber side, in agreement with the substitution of Sr\(^{2+}\) sites with \(C_3\) point symmetry. In addition, it is reported that the intensity ratio of \(5D_0\)-\(F_2\) transition to \(5D_0\)-\(F_2\) transition commonly gives information on the magnitude of the J-mixing effect associated with the \(5D_0\)-\(F_2\) transition.\(^{32}\) The very small ratio \((0.013)\) in the present case suggests that the J-mixing effect can be neglected.

Fig. 8 displays the vacuum referred binding energy (VRBE) scheme for all lanthanide 4f (bar) and 5d (circle) states in KSrPO\(_4\). Curve a connects the lowest 5d state of trivalent lanthanide and curve b that of the lowest 5d state of divalent lanthanide.
To construct this VRBE scheme, the Coulomb repulsion energy \( U \), which is the energy difference between the ground state 4f-shell electron binding energies \( E_{4f}(Eu^{3+}) \) in \( Eu^{3+} \) and \( E_{4f}(Eu^{2+}) \) in \( Eu^{2+} \), is first estimated to be about 7.09 eV from the observed Ce\(^{3+} \) 5d centroid shift \( \varepsilon_c \) (1.19 eV) using Eq. (1), then the energy of \( Eu^{3+} \) ground state \( 8S_{7/2} \) in the VRBE scheme is evaluated to be -4.11 eV based on Eq. (2).

\[
U = 5.44 + 2.834e^{-c/2} \tag{1}
\]

\[
E_{4f}(Eu^{2+}) = -24.92 + \frac{18.05 - U}{0.777 - 0.0353U} \tag{2}
\]

Therefore, the \( Eu^{2+} \) ground state \( 7F_0 \) is about -11.20 eV in VRBE. The 4f ground state energies of other Ln\(^{3+} \) and Ln\(^{2+} \) are determined using the most recently reported energy differences of \( Eu^{3+} / Eu^{2+} \) and \( Ln^{3+} / Ln^{2+} \) in refence. Then all 4f energies of Ln\(^{3+} \) and Ln\(^{2+} \) can be drawn in the VRBE scheme according to Dieke diagram.

After we obtain the 4f ground state \((F_{5/2}) \) energy of \( Ce^{3+} \) about -5.96 eV in VRBE scheme, the \( Ce^{3+} \) 5d state can be derived to be approximately -1.82 eV by adding the observed energy of 4f-5d\(_1 \) transition (4.14 eV, see in Fig. 3). The lowest 5d energies of other Ln\(^{3+} \) can be determined with the known energy differences of \( Ce^{3+} \) and \( Ln^{3+} \).

The 4f-5d\(_1 \) transition energy of \( Eu^{2+} \) (~3.16 eV) is measured experimentally as show in Fig. 10 or estimated from that of \( Ce^{3+} \). Hence the 5d\(_1 \) energy of \( Eu^{2+} \) in VRBE scheme is determined to be -0.95 eV and the 5d\(_i \) energies of other Ln\(^{3+} \) are obtained from the energy differences of \( Eu^{2+} \) and \( Ln^{3+} \).

The top of valence band energy of \( KSrPO_4 \) in VRBE scheme is estimated be -8.67 eV by subtracting \( Eu^{3+} \)\( \rightarrow \) charge transfer energy (4.56 eV) from \( E_{4f}(Eu^{2+}) \) (-4.11 eV). Subsequently, the bottom of conduction band (CB) energy of \( KSrPO_4 \) in VRBE scheme is -0.55 eV after adding the band gap energy (8.12 eV) of \( KSrPO_4 \).

Based on the VRBE scheme, the energy difference between the energy level of \( Ce^{3+} \) 5d\(_1 \) state and the bottom of conduction band is approximately 1.27 eV. The activation energy \( \Delta E \) of thermal quenching of \( Ce^{3+} \) 5d-4f luminescence generally corresponds to the energy required to raise an electron from the relaxed 5d\(_i \) excited level into the host conduction band. This large energy gap explains that luminescence of \( Ce^{3+} \) doped \( KSrPO_4 \) phosphor possesses a good thermal stability.

3.4. The f-d Transitions of \( Pr^{3+} \) in \( KSrPO_4 \)

Fig. 9 shows the synchrotron radiation VUV-UV excitation spectrum (\( \lambda_{ex}=269 \) nm) of sample \( K_{0.003}Sr_{0.994}Pr_{0.003}PO_4 \) at 25 K; the inset shows the emission spectrum (\( \lambda_{ex}=213 \) nm) at 25 K and the corresponding Gaussian fitting results.

3.5. The Redox Properties of \( Eu \) in \( K SrPO_4 \)
Fig. 10 displays the synchrotron radiation VUV-UV excitation (curve a) and laboratory UV-vis excitation (curve b) and emission spectra of sample K\textsubscript{Sr}0.999\textsubscript{Eu}0.001\textsubscript{PO}_4 prepared in CO ambiance. Considering that in the 370-470 nm range there is no Eu\textsuperscript{3+} emission as shown in Fig. 10 (b), the excitation profile contains only the weak host-related absorption in short-wavelength range and the 4f-5d transitions of Eu\textsuperscript{2+} in the long-wavelength range when the emission wavelength of 422 nm is monitored. The lowest 4f\textsuperscript{7}-4f\textsuperscript{6}5d excitation band in our case is about 392 nm (3.16 eV), which is empirically estimated as the energy at 15% to 20% of the maximum excitation intensity on the long-wavelength side. This is close to the ZPLs position (~396 nm) as shown in ESI Fig. S7\textsuperscript{†}. Referring to the VRBE scheme (Fig. 8), the energy barrier for electrons from excited state of Eu\textsuperscript{2+} to the bottom of conduction band (\(\Delta E\)) is estimated to be 0.40 eV. A small \(\Delta E\) energy usually makes the Eu\textsuperscript{2+} emission prone to thermal quenching by means of the ionization of the 5d electrons. The large \(\Delta E\) value in present case implies a good thermal stability of luminescence of Eu\textsuperscript{2+} in K\textsubscript{Sr}PO\textsubscript{4}. The temperature-dependent decay dynamics in Fig. 11 confirms this standpoint, the thermal quenching of Eu\textsuperscript{2+} luminescence does not occur until 500 K.

Fig. 11 Temperature dependent decay curves (\(\lambda_\text{ex} = 325\) nm, \(\lambda_\text{em} = 422\) nm) of K\textsubscript{Sr}0.999\textsubscript{Eu}0.001\textsubscript{PO}_4 sample prepared in CO ambiance.

Fig. 12 show emission spectra of K\textsubscript{Sr}0.999\textsubscript{Eu}0.001\textsubscript{PO}_4 prepared in CO reducing, N\textsubscript{2} inert and air oxidizing conditions. We selected 319 nm as excitation wavelength because both Eu\textsuperscript{3+} and Eu\textsuperscript{2+} absorb at this wavelength [see Fig. 6 and 10 (a)]. In Fig. 12, the broad band at 425 nm belongs to d-f emission of Eu\textsuperscript{2+} and several sharp lines in long-wavelength correspond to Eu\textsuperscript{3+} f-f emission. Apparently, Eu\textsuperscript{3+} can be reduced to Eu\textsuperscript{2+} in CO reducing, N\textsubscript{2} inert, and air oxidizing conditions, but the reduction degrees are different. To express the degree of reduction of Eu\textsuperscript{3+} at these conditions, we calculated the ratios (\(A_2\)/\(A_3\)) of the integrated area of emission between 380-450 (\(A_2\)) and between 550-730 (\(A_3\)) nm, respectively. The results are 99.8/0.02 (CO), 46.4/53.6 (N\textsubscript{2}), 31.9/68.1 (air), implying that the reduction degree decreases from CO to N\textsubscript{2} and air. When prepared in CO condition, it dominantly displays Eu\textsuperscript{2+} emission, and Eu\textsuperscript{3+} emission is very weak. For further demonstration, under excitation in the charge transfer band of Eu\textsuperscript{3+} (272 nm), the emission spectra of representative K\textsubscript{Sr}x\textsubscript{Eu}y\textsubscript{PO}_4 (x = 0.001, 0.009 and 0.015) samples annealing in CO conditions.
condition are shown in ESI Fig. S8†, indicating Eu\(^{2+}\) ions are dominant. In N\(_2\) and air conditions, Eu\(^{2+}\) emission dramatically decreases but is still present, showing that Eu\(^{3+}\) is partially reduced to Eu\(^{2+}\), and it seems that more Eu\(^{3+}\) ions are reduced to Eu\(^{2+}\) when prepared under N\(_2\) conditions. The redox behaviour of Eu in KSr\(_{0.995}\)Eu\(_{0.005}\)PO\(_4\) can be understood with the VRBE scheme in Fig. 8 when we consider that the redox properties of Eu ions are related to the Fermi energy position. The Fermi energy level locates midway the bottom of the CB and the top of the valence band (VB) when the phosphor is synthesized in inert atmosphere or near -4.06 eV (black dot curve) in the VRBE scheme. Eu ions usually prefer the divalent state when the ground state of Eu\(^{3+}\) is near or below the Fermi energy, and becomes more stable when it is further below the Fermi energy level. Referring to VRBE scheme, it is almost same as the Fermi energy level for KSr\(_{0.995}\)Eu\(_{0.005}\)PO\(_4\) sample annealed in N\(_2\), so Eu\(^{3+}\) can be partially reduced to Eu\(^{2+}\) in N\(_2\). Fermi energy position somewhat increase or decrease at oxidizing or reducing conditions in comparison with that at inert condition. As a result, the reduction degree of Eu\(^{3+}\) increases in CO reducing condition but decreases in air oxidizing condition. On the other hand, when we consider the structure of the host compound, the stabilization of Eu\(^{2+}\) would relate to the similar ionic radii of Sr\(^{2+}\) and Eu\(^{2+}\), and the rigid tetrahedral three-dimensional PO\(_4^{3-}\) groups as previous comments.\(^{5,37}\)

4. Conclusions

We investigated the synchrotron radiation VUV-UV excitation spectra and corresponding emission spectra of Ln-doped KSrPO\(_4\) (Ln = Ce\(^{3+}\), Eu\(^{3+}\), Eu\(^{2+}\), Pr\(^{3+}\)) phosphors prepared by a high-temperature solid-state reaction method. It is found that five 4f-5d transitions of Ce\(^{3+}\) are at ~4.14, 4.60, 5.23, 5.76 and 6.08 eV, respectively. The energy of the Eu\(^{2+}\)-O\(^2-\) charge transfer state is found at 4.56 eV (272 nm), and the mobility band gap of KsrPO\(_4\) is estimated to be 8.12 eV. Accordingly, the VRBE scheme is constructed and applied to understand the excellent thermally stable luminescence of Ce\(^{3+}\) and Eu\(^{2+}\), 4f-5d transitions of Pr\(^{3+}\), and redox properties of Eu\(^{2+}/3+\).

The energy differences between the lowest 5d states of Ce\(^{3+}\) / Eu\(^{3+}\) and the bottom of the conduction band are read about 1.27 / 0.40 eV from VRBE scheme, individually. Referring to the experimental data and derived energy barriers, the 5d-4f luminescence of Ce\(^{3+}\) and Eu\(^{3+}\) has a good thermal stability and the thermal quenching does not occur as far as 500 K. In terms of the VRBE scheme, the lowest 5d energy of Pr\(^{3+}\) is evaluated to be about 5.64 eV, which is in good line with the experimental observation (5.66 eV), and other four 5d orbital of Pr\(^{3+}\) are with energies approximately 6.12, 6.75, 7.28, and 7.60 eV, respectively. Eu\(^{3+}\) can be reduced to Eu\(^{2+}\) in CO reducing, N\(_2\) inert, and air oxidizing conditions, respectively. The reduction is almost totally in CO reducing ambiance, while that is partially in N\(_2\) inert and air oxidizing atmosphere. From the viewpoint of host structure, the similar ionic radii of Sr\(^{2+}\) and Eu\(^{2+}\), and the rigid tetrahedral three-dimensional PO\(_4^{3-}\) groups may be responsible for the phenomenon. From the standpoint of energy, this may relate to nearly same energies of Eu\(^{2+}\) ground state and the Fermi level of KsrPO\(_4\) in neutral condition.

**Conflicts of interest**

There are no conflicts to declare.

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**Notes and references**

33 P. Dorenbos, J. Lumin., 2005, 111, 89.